

Acids & Bases HOMEWORK

Lesson 1

1. Provide the name for each of the strong acids listed below. Provide the formula for each of the strong bases listed below.

- ~~HBr~~ **hydrobromic acid**
- ~~HCl~~ **hydrochloric acid**
- ~~HNO₃~~ **nitric acid**
- ~~H₂SO₄~~ **sulfuric acid**
- ~~Strontium hydroxide~~ **LiOH**
- ~~potassium hydroxide~~ **KOH**
- ~~calcium hydroxide~~ **CaOH**
- ~~calcium hydroxide~~ **Ca(OH)₂**
- ~~ammonium hydroxide~~ **Li(OH)₄**

~~H₂O₂ for HYDROPEROXIDE acids it never will completely dissociate in water~~

2. Of the acids and bases listed below, circle the ones that completely dissociate in water.

HCl

~~HF~~

~~H₂SO₄~~

~~Mg(OH)₂~~

NaOH

NaHCO₃

~~NaOH~~

Na₂CO₃

3. What is the formula for hydronium? **H₃O⁺**

4. BaC_2O_4 will react with water and lose a H^+ ion in the process.

- Write a balanced equation for this reaction: $\text{BaC}_2\text{O}_4 + \text{H}_2\text{O} \rightleftharpoons \text{BaC}_2\text{O}_4^- + \text{HC}_2\text{O}_4^-$
- Would you expect a large or small equilibrium constant for this reaction? **small, because BaC_2O_4 is a weak acid.**
- Would you define BaC_2O_4 as a proton donor or a proton acceptor? **proton donor, because it is losing a H^+ ion.**
- In your written equation, is BaC_2O_4 acting as a Brønsted-Lowry acid or a Brønsted-Lowry base? **Brønsted-Lowry acid, because it's a proton donor.**
- Identify the conjugate base of BaC_2O_4 : **BC_2O_4^-**
- Identify BaC_2O_4 as polyprotic, amphoteric, or neither: **polyprotic, because it contains multiple H's to give.**

5. F^- will react with water and gain a H^+ ion in the process.

- Write a balanced equation for this reaction: $\text{F}^- + \text{H}_2\text{O} \rightleftharpoons \text{HF} + \text{OH}^-$
- Would you expect a large or small equilibrium constant for this reaction? **small, because F^- is a weak base.**
- Would you define F^- as a proton donor or a proton acceptor? **proton acceptor, because it is gaining a H^+ ion.**
- In your written equation, is F^- acting as a Brønsted-Lowry acid or a Brønsted-Lowry base? **Brønsted-Lowry base, because it's a proton acceptor.**
- Identify the conjugate acid of F^- : **HF**
- Identify F^- as polyprotic, amphoteric, or neither: **neither**

6. HSO_4^- will react with water and lose a H^+ ion in the process.

- Write a balanced equation for this reaction: $\text{HSO}_4^- + \text{H}_2\text{O} \rightleftharpoons \text{SO}_4^{2-} + \text{H}_3\text{O}^+$
- Would you expect a large or small equilibrium constant for this reaction? **small, because HSO_4^- is a weak acid.**
- Would you define HSO_4^- as a proton donor or a proton acceptor? **proton donor, because it is losing a H^+ ion.**
- In your written equation, is HSO_4^- acting as a Brønsted-Lowry acid or a Brønsted-Lowry base? **Brønsted-Lowry acid, because it's a proton donor.**
- Identify the conjugate base of HSO_4^- : **SO_4^{2-}**
- Identify HSO_4^- as polyprotic, amphoteric, or neither: **amphoteric, because it has a negative charge (even with H's it hasn't lost all the H's) and it has H's to give.**

19 3 Acids And Bases Answer Key Chemistry

Henry Watts



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Cumulated Index Medicus ,1995 Electrochemistry in Nonaqueous Solutions Kosuke Izutsu, 2009-09-22 An excellent resource for all graduate students and researchers using electrochemical techniques After introducing the reader to the fundamentals the book focuses on the latest developments in the techniques and applications in this field This second edition contains new material on environmentally friendly solvents such as room temperature ionic liquids

High Temperature Polymer Electrolyte Membrane Fuel Cells Qingfeng Li, David Aili, Hans Aage Hjuler, Jens Oluf Jensen, 2015-10-15 This book is a comprehensive review of high temperature polymer electrolyte membrane fuel cells PEMFCs PEMFCs are the preferred fuel cells for a variety of applications such as automobiles cogeneration of heat and power units emergency power and portable electronics The first 5 chapters of the book describe rationalization and illustration of approaches to high temperature PEM systems Chapters 6-13 are devoted to fabrication optimization and characterization of phosphoric acid doped polybenzimidazole membranes the very first electrolyte system that has demonstrated the concept of and motivated extensive research activity in the field The last 11 chapters summarize the state of the art of technological development of high temperature PEMFCs based on acid doped PBI membranes including catalysts electrodes MEAs bipolar plates modelling stacking diagnostics and applications

The American chemist ,1874

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