

14.3

HYDROGEN IONS AND ACIDITY

Section Review

Objectives

1. Calculate concentrations based on the pH of acids and bases and vice versa.
2. Calculate pH and concentrations based on K_a and K_b and vice versa.
3. Calculate the pH of weak acids and bases.

Vocabulary

- | | | |
|------------------------|-----------------------------|-------------------|
| 1. autoionization | 2. conjugate acid-base pair | 3. weak acid/base |
| 4. ionization | 5. K_a | 6. K_b |
| 7. ionization constant | 8. pH | |

Key Equation

- | | |
|-----------------------|-----------------------|
| 1. $[H^+] = 10^{-pH}$ | 2. $pH = -\log [H^+]$ |
| 3. $pH = -\log [H^+]$ | 4. $[H^+] = 10^{-pH}$ |

Part A Completion

Write the equation and/or the numerical value(s) of the constant(s) that are indicated in the spaces provided. Be sure to include proper units when appropriate.

1. Water autoionizes to H^+ and hydroxide ions. The concentration of H^+ in the autoionization of water at 25°C is 1.0×10^{-7} mol/L.
2. The autoionization constant for water is 1.0×10^{-14} at 25°C.
3. The K_a of acetic acid is 1.8×10^{-5} .
4. The K_b of ammonia is 1.8×10^{-5} .
5. The K_a of hydrofluoric acid is 6.6×10^{-4} .
6. The K_b of the conjugate base of acetic acid is 5.6×10^{-10} .
7. The K_a of the conjugate acid of ammonia is 5.6×10^{-10} .

8. K_a of HCl
9. K_b of OH^-
10. K_a of H_2O
11. K_b of H_2O
12. K_a of $HClO_4$
13. K_b of ClO_4^-
14. K_a of HNO_3
15. K_b of NO_3^-
16. K_a of H_2SO_4
17. K_b of HSO_4^-
18. K_a of H_2SO_4

19 2 Hydrogen Ions And Acidity Work Sheet Answers

Wilhelm Mansfield Clark



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