

# Data:

## Experiment 6: The Atomic Spectrum of Hydrogen

### 1. Calculate the Values of the Energy Levels of the Hydrogen Atom

| n | Energy, $E_n$                     | n | Energy, $E_n$                     |
|---|-----------------------------------|---|-----------------------------------|
| 1 | $-2.18 \times 10^{-18} \text{ J}$ | 4 | $-1.36 \times 10^{-19} \text{ J}$ |
| 2 | $-5.45 \times 10^{-19} \text{ J}$ | 5 | $-8.72 \times 10^{-20} \text{ J}$ |
| 3 | $-2.42 \times 10^{-19} \text{ J}$ | 6 | $-6.05 \times 10^{-20} \text{ J}$ |

### 2. Calculation of the Wavelengths of the Lines in the Hydrogen Spectrum

In the upper half of each box write the difference in energy between  $E_{\text{obs}}$  and  $E_{\text{ref}}$ . In the lower half of the box write the wavelength associated with that value.

| $n_{\text{hi}} \downarrow$<br>$n_{\text{lo}} \rightarrow$ | 6                                            | 5                                            | 4                                           | 3                                          | 2                                          |
|-----------------------------------------------------------|----------------------------------------------|----------------------------------------------|---------------------------------------------|--------------------------------------------|--------------------------------------------|
| 1                                                         | $2.12 \times 10^{-17} \text{ J}$<br>94.6 nm  | $2.09 \times 10^{-17} \text{ J}$<br>95.0 nm  | $2.04 \times 10^{-18} \text{ J}$<br>97.3 nm | $1.94 \times 10^{-19} \text{ J}$<br>103 nm | $1.63 \times 10^{-19} \text{ J}$<br>122 nm |
| 2                                                         | $4.845 \times 10^{-19} \text{ J}$<br>411 nm  | $4.578 \times 10^{-19} \text{ J}$<br>434 nm  | $4.09 \times 10^{-19} \text{ J}$<br>486 nm  | $3.03 \times 10^{-20} \text{ J}$<br>656 nm |                                            |
| 3                                                         | $1.815 \times 10^{-19} \text{ J}$<br>1095 nm | $1.548 \times 10^{-19} \text{ J}$<br>1282 nm | $1.06 \times 10^{-19} \text{ J}$<br>1875 nm |                                            |                                            |
| 4                                                         | $7.55 \times 10^{-20} \text{ J}$<br>2633 nm  | $4.88 \times 10^{-20} \text{ J}$<br>4073 nm  |                                             |                                            |                                            |
| 5                                                         | $2.67 \times 10^{-20} \text{ J}$<br>7445 nm  |                                              |                                             |                                            |                                            |

### 3. Assignment of Lines in the Hydrogen Spectrum

# Atomic Spectrum Of Hydrogen Pre Lab Answers

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