

CHEMISTRY ADVANCED STUDY

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FUNDAMENTAL CONCEPTS

Math

- Properties of exponents
- Properties of logarithms
 - $\log_b (M^p N) = \log_b (M) + \log_b (N)$
 - $\log_b (MN) = \log_b (M) + \log_b (N)$
 - $\log_b (M^p) = p \log_b (M)$
- Graphing slope intercepts
 - $y = mx + b$
 - b is the y-intercept, m is the slope
- Quadratic formula
 - $x = \frac{-b \pm \sqrt{b^2 - 4ac}}{2a}$
 - Ex. $x^2 + 4x - 21$
 - $x = \frac{-4 \pm \sqrt{16 - 4(1)(-21)}}{2}$
 - $x = -2 \pm 5$
 - $x = 3$ or $x = -7$

Science

- Scientific notation
- Dimensional analysis
- Significant figures

Physics and Biology

- Scientific method
- Coulomb's law: electric force between like or opposite charges; dependence on separation distance and charge magnitude

ELEMENTS AND ATOMS

Protons: define the element (atomic number), positive

Electrons: define the charge (bonds and polarity), negative

Neutrons: define isotopes, neutral

ATOMIC MASS AND WEIGHT

Atomic Mass: Carbon-12 (exactly 12 atomic mass unit)

Atomic Weight

- weighted average of atomic masses of element
- (atomic mass · neutrons) + (atomic mass · neutrons)

MOLE A.K.A. AVOGADRO'S NUMBER

$$= 6.02 \cdot 10^{23}$$

In Terms of Carbon

1 mole of ^{12}C atoms in 12g of Carbon

$6.02 \cdot 10^{23} = ^{12}\text{C}$ atoms in 12g of Carbon

Gram to AMU

1 gram = 1 mole of atomic mass units

1 gram = $6.02 \cdot 10^{23}$ amu

1 mol ^{27}Al = 27g

2 mol ^{28}Si = 56g

4 mol ^{16}O = 64g

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