

# Atomic Structure Review

Name \_\_\_\_\_

## 1. The 3 particles of the atom are:

	Location in Atom?	Size of LOCATION?	Charge?	Mass?
Protons	nucleus	.0001 %	+	1 amu
Neutrons	nucleus	.0001 %	neutral	1 amu
Electrons	cloud	99.999 %	-	.00054 amu

2. The number of protons in one atom of an element determines the atom's <sup>properties</sup> ~~charge~~ ~~#~~, and the number of electrons determines the charge of an element.

3. The atomic number tells you the number of protons in one atom of an element. It also tells you the number of electrons in a neutral atom of that element.

4. The atomic number gives the "identity" of an element as well as its location on the Periodic Table. No two different elements will have the same atomic number.

5. The avg atomic mass of an element is the average mass of an element's naturally occurring atom, or isotopes, taking into account the # of neutrons of each isotope.

6. The mass # of an element is the total number of protons and neutrons in the nucleus of the atom.

7. The mass number is used to calculate the number of neutrons in one atom of an element. In order to calculate the number of neutrons you must subtract the protons from the neutrons.

# Atomic Structure Guide Answers

**Mike Jess**

A red circular graphic with a gradient, appearing as a semi-circle or a partial circle, located to the right of the author's name.

## **Atomic Structure Guide Answers:**

**Atom Structure Guide** Sophia Anderson, AI, 2025-02-22 Atom Structure Guide explores the core concepts of atomic physics providing a comprehensive look at the atom's structure and behavior through the lens of quantum mechanics and experimental findings. Consider that the quantum mechanical model replaced older classical notions with probabilities revolutionizing our understanding of chemical bonding and material properties. Further experimental techniques like spectroscopy and electron microscopy have offered crucial insights into the atom's internal structure. This book uniquely balances theoretical concepts like wave-particle duality and the Schrödinger equation with experimental evidence like spectroscopic measurements and microscopic images. It emphasizes how a solid grasp of atomic structure is essential for progress across scientific fields. The book starts with the historical development of atomic theory and systematically builds upon that foundation dedicating chapters to various spectroscopic techniques and concluding with the atom's role in chemical bonding, material properties, and quantum computing. The book's approach is designed to make complex topics accessible without sacrificing scientific accuracy making it a valuable resource for students, researchers, and anyone interested in the fundamentals of physics and science.

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