

EQ: Describe atomic structure.

- Atoms are made of extremely tiny particles called protons, neutrons, and electrons.
- Protons and neutrons are in the center of the atom, making up the nucleus.
- Electrons surround the nucleus.
- Protons have a positive charge.
- Electrons have a negative charge.
- The charge on the proton and electron are exactly the same size but opposite.
- Neutrons have no charge.
- Since opposite charges attract protons and electrons attract each other.

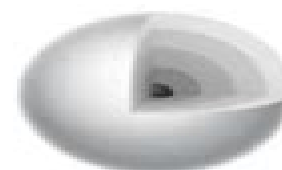
- ☐ Label the nucleus (protons, neutrons) and electrons in the drawing of a carbon atom.
- ☐ Draw a line between the subatomic particle and its charge.

proton
electron
neutron

no charge
positive charge
negative charge



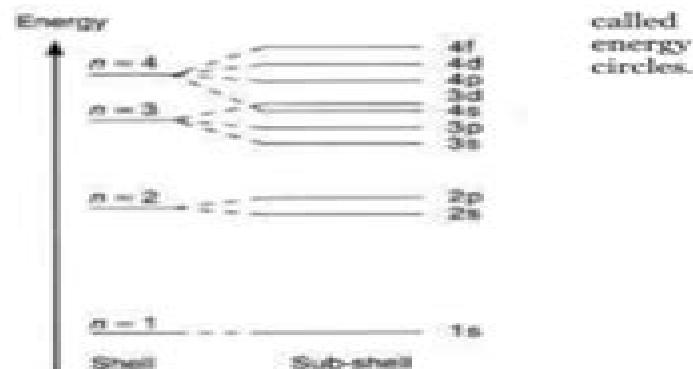
- The electrons surrounding an atom are located in regions around the nucleus called "energy levels".
- An energy level represents the 3-dimensional space surrounding the nucleus where electrons are most likely to be.
- The first energy level is closest to the nucleus. The second energy level is a little farther away than the first. The third is a little farther away than the second, and so on.
- Each energy level can accommodate or "hold" a different number of electrons before additional electrons begin to go into the next level.
- When the first energy level has 2 electrons, the next electrons go into the second energy level until the second level has 8 electrons.
- When the second energy level has 8 electrons, the next electrons go into the third energy level until the third level has 8 electrons.
- When the third energy level has 8 electrons, the next 2 electrons go into the fourth energy level.
- The electrons in the energy level farthest from the nucleus are called valence electrons.
- Atoms in the same column (group) in the periodic table have the same number of valence electrons.



Energy levels

Electrons surround the nucleus of an atom in regions *energy levels*. Even though atoms are spherical, the levels in an atom are more easily shown in concentric

Energy Level	Maximum # electrons
1	2
2	8
3	18
4	32
5	50
6	72



Atomic Structure Study Guide

James E. Brady, John R. Holum



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