Atoms & The Periodic Table Unit 3 Test Study Guide

Lesson 1: The Atom

- 1. Who was the ancient Greek philosopher that came up with the idea that all things are made of small "uncuttable" particles?
- 2. What is an atom?
- 3. What was John Dalton's contribution to the atomic theory?
- 4. What are the parts of an atom?
 - a. Where can each of these particles be found within an atom?
 - b. How do the electrical charges of protons, neutrons, and electrons compare to each other?
 - c. How do the masses of protons, neutrons, and electrons compare to each other?
- 5. What is an atomic number?
 - a. What does this tell about the structure of an atom?
- 6. What is the atomic mass (mass number)?
 - a. What does this tell about the structure of an atom?
 - i. How can the atomic mass help you find the number of neutrons in an atom?

Lesson 2: The Periodic Table

- 1. What are elements?
- 2. What is the Periodic Table of Elements?
 - a. How are elements arranged on the Periodic Table?
- 3. Information found on each square of the Periodic Table:
 - a. Atomic Number (See Lesson 1)
 - b. What is a chemical name?
 - i. Examples
 - c. What is a chemical symbol?
 - Examples
 - d. Atomic Mass (See Lesson 1)
- 4. What are the 3 types of elements?
 - a. Where can these elements be found on the Periodic Table?
 - b. What are the characteristics of these types of elements?
- 5. What is a group?
 - a. What do the elements in a group have in common?
- 6. What is a period?
 - a. What do the elements in a period have in common?

Lesson 3: Electrons and Chemical Bonding

- 1. What is a chemical bond?
 - a. Why do atoms form bonds?
 - i. How can you figure out what an atom will do to become "happy"?
- 2. What is a group of atoms that are chemically bonded called?
- 3. What are valence electrons?

Atoms And Periodic Table Study Guide Answer

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