

Name _____ Section _____

Acid-Base Titration (Volumetric Analysis) – Data Sheet

CHEM 161A – General Chemistry II Lab

Attach all calculations as separate pages.

Part C – Standardization of NaOH

$$M_a V_a = M_b V_b$$

Standardization of NaOH	Trial 1	Trial 2
Mass $KC_8H_5O_4$ (KHP) weighed (g)	0.4275(g)	0.4275(g)
Molar mass $KC_8H_5O_4$ (KHP) (g/mol)	204.22121(g/mol)	204.22121(g/mol)
Moles $KC_8H_5O_4$ (KHP) used (mol) (MaVa)	0.002093 mol	0.002093 mol
Stoichiometric ratio of KHP to NaOH	1 mol KHP:1 mol NaOH	1 mol KHP:1 mol NaOH
Moles NaOH neutralized (mol)	+	
Initial volume reading NaOH from buret (mL)	20.01 mL	20.01 mL
Final volume reading NaOH from buret (mL)	37.26 mL	37.26 mL
Volume NaOH used (mL) (V)	17.25 mL	17.25 mL
Molarity of NaOH (mol/L)		
Average molarity of NaOH (mol/L)		

= ?
2

Part D – Titration of an Unknown Diprotic Acid

Molar Mass of an Unknown Acid	Trial 1	Trial 2
Mass of unknown acid (g)	0.1072 g	0.1244 g
Standardized molarity of NaOH (mol/L) [Part C]		
Initial volume reading NaOH from buret (mL)	0.31 mL	0.25 mL
Final volume reading NaOH from buret (mL)	18.53 mL	21.72 mL
Volume NaOH used (mL)		
Moles NaOH used (mol)		
Stoichiometric ratio of NaOH to acid		
Moles of acid neutralized (mol)		
Molar mass unknown acid (g/mol)		
Average molar mass unknown acid (g/mol)		

(Answer)

Acid Base Titration And Volumetric Analysis Lab 42

Answers

Thomas Griffiths



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