

Lesson 14.2 Acid and Base Strengths

Suggested Reading

- Zumdahl Chapter 14 Section 14.1, 14.2, 14.6, & 14.7

Essential Question

- What is the relationship between the strength of an acid (base) the the extent of dissociation?



Learning Objectives:.

- Distinguish between strong and weak acids and bases in terms of the extent of dissociation, reaction with water and electrical conductivity.
- State whether a given acid or base is strong or weak.
- Distinguish between strong and weak acids and bases, and determine the relative strengths of acids and bases using experimental data.
- Write the equilibrium expression for acid dissociation in water.

Introduction

The Bronsted-Lowry concept considers an acid-base reaction as a proton-transfer reaction. It is useful to consider acid-base reactions as competitions between species for protons. From this point of view, you can order acids by their relative strengths as proton donors. The strong acids are those that lose their protons more easily than other acids. Similarly, the stronger bases are the ones that hold on to protons more strongly than other bases. Lets look at objective one from above.

Distinguish between strong and weak acids and bases in terms of the extent of dissociation, reaction with water and electrical conductivity:

<u>Strong Acid</u>	<u>Weak Acid</u>	<u>Strong Base</u>	<u>Weak Base</u>
Completely dissociates	Partly dissociates	Completely dissociates	Partly dissociates

Acids And Bases Study Guide

**T. W. Graham Solomons, Craig B.
Fryhle, Scott A. Snyder**



Acids And Bases Study Guide:

Chemistry, Student Study Guide James E. Brady, Fred Senese, 2008-01-28 The image on the front cover depicts a carbon nanotube emerging from a glowing plasma of hydrogen and carbon as it forms around particles of a metal catalyst Carbon nanotubes are a recently discovered allotrope of carbon Three other allotropes of carbon buckyballs graphite and diamond are illustrated at the left as is the molecule methane CH_4 from which nanotubes and buckyballs can be made The element carbon forms an amazing number of compounds with structures that follow from simple methane found in natural gas to the complex macromolecules that serve as the basis of life on our planet The study of chemistry also follows from the simple to the more complex and the strength of this text is that it enables students with varied backgrounds to proceed together to significant levels of achievement

Basic Concepts of Chemistry, 9e Study Guide and Solutions Manual Leo J. Malone, Theodore O. Dolter, 2012-01-03 The 9th edition of Malone's Basic Concepts of Chemistry provides many new and advanced features that continue to address general chemistry topics with an emphasis on outcomes assessment New and advanced features include an objectives grid at the end of each chapter which ties the objectives to examples within the sections assessment exercises at the end each section and relevant chapter problems at the end of each chapter A new Math Check allows quick access to the needed basic skill The first chapter now includes brief introductions to several fundamental chemical concepts and Chapter Synthesis Problems have been added to the end of each chapter to bring key concepts into one encompassing problem Every concept in the text is clearly illustrated with one or more step by step examples Making it Real essays have been updated to present timely and engaging real world applications emphasizing the relevance of the material they are learning This edition continues the end of chapter Student Workshop activities to cater to the many different learning styles and to engage users in the practical aspect of the material discussed in the chapter

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Organic Chemistry, 12e Binder Ready Version Study Guide & Student Solutions Manual T. W. Graham Solomons,Craig B. Fryhle,Scott A. Snyder,2016-04-11 This is the Student Study Guide Solutions Manual to accompany Organic Chemistry 12th Edition The 12th edition of Organic Chemistry continues Solomons Fryhle Snyder s tradition of excellence in teaching and

preparing students for success in the organic classroom and beyond A central theme of the authors approach to organic chemistry is to emphasize the relationship between structure and reactivity To accomplish this the content is organized in a way that combines the most useful features of a functional group approach with one largely based on reaction mechanisms The authors philosophy is to emphasize mechanisms and their common aspects as often as possible and at the same time use the unifying features of functional groups as the basis for most chapters The structural aspects of the authors approach show students what organic chemistry is Mechanistic aspects of their approach show students how it works And wherever an opportunity arises the authors show students what it does in living systems and the physical world around us

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