Model 2 – Ternary Acids (Oxyacids)

Acid	Name of Acid in Aqueous Solution	Cation (+)	Polyatomic Anion	Polyatomic Anion Name
HCIO.	Chloric acid	H,O-	C/O3-1	Chlorate
H,SO,	Sulfarous acid	2450	503-2	Sulfite
H,SO,	Sulfuric acid	2150	500-3	Sulfate
H.PO,	Phosphorous acid	3 H,O-	P03-3	Phosphite
H.PO.	Phosphoric acid	3 H50+	PO,5-	Phosphate
HNO,	Nitric acid	H30+	NO."	Nitnate
HNO,	Nisrous acid	H.O+		Nitrite
H,CO.	Carbonic acid	2 H30+	CO,-2	Carbonate

- 6. Look at the formulas of the ternary acids in Model 2.
 - at. How are ternary acids different from binary acids in their structure?

Ternory acids contain three elements and they also contain the element oxygen.

6. What number do you think the prefix "ter-" refers to?

"te" refer to three

- 7. When ternary acids are mixed with water, ions will form. Fill in the table above with the formulas and names of the anions.
- 8. Examine the pairs of ternary acids in Model 2 that contain sulfur, phosphorus, and nitrogen. Each pair has one acid that ends in "-ic" and another that ends in "-ous." These endings are related to the name of the polyatomic anion found in the acid ("-ate" or "-ite"). Complete the statements below with the correct acid name ending.

Polyatomic anion ending is "-ate" → acid name ending is /C.

- Polyatomic anion ending is "-ite" → acid name ending is aus
- 9. If the prefix "hydro-" were used to name a ternary acid, what problem would this create when
- naming HClO? C/O3' has the name of Chlorate. "ate" must be replaced with "ic. If "hydro" is used as a prefix, it would be hydrochloric acid. This has the formula HCl, not HClO3.

 The prefix "hydro" is not used in the name when an acid contain more than two elements and are of the elements is axygen.

 The prefix "hydro" is not used in the name when an acid contain more than two elements and are of the elements is axygen.

 It is replaced with it and it is replaced with "ous".

12. Circle the acid(s) below that would be named beginning with the prefix "hydro-."

HISO.



H.CO.

HNO.

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NARAYAN CHANGDER

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most widely quoted books in the chemical literature Although ultimately there is no satisfactory alternative to experimental measurement it is not always convenient or practicable to make the necessary measure ments and calculations Moreover the massive pK compilations currently available provide values for only a small fraction of known or possible acids or bases For example the compilations listed in Section 1 3 give pK data for some 6 000 8 000 acids whereas if the conservative estimate is made that there are one hundred different substituent groups available to substitute in the benzene ring of benzoic acid approximately five million tri substituted benzoic acids are theoretically possible Thus we have long felt that it is useful to consider methods by which a pK value might be predicted as an interim value to within several tenths of a pH unit using arguments based on linear free energy relationships by analogy by extrapolation by interpolation from existing data or in some other way This degree of precision may be adequate for many purposes such as the recording of spectra of pure species as anion neutral molecule or cation for selection of conditions favourable to solvent extraction and for the interpretation of Acids, Bases, and Salts Brian J. Knapp,1998 pH profiles for organic reactions More Acids and Bases ,1944 Acids Acids and Bases Norris Folger Hall,1941 and Bases ,2010 Acids and Bases Ray T. Slattery, 1973 **Hard and Soft** Acids and Bases ,1973

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