Data:

Experiment 6; The Atomic Spectrum of Hydrogen

1. Calculate the Values of the Energy Levels of the Hydrogen Atom

n	Energy, E _n	n Energy, E _n	
1	-2.18×10-14	4	C "-01×00.1-
2	L^1-01 × 24.2 -	5	C 05-01 × 25.8-
3	- 2.42 × 10-19J	6	-6.05 × 10-30 J

2. Calculation of the Wavelengths of the Lines in the Hydrogen Spectrum

In the upper half of each box write the difference in energy between E_{nbi} and E_{nbo} . In the lower half of the box write the wavelength associated with that value.

$n_{lo} \downarrow n_{hi} \rightarrow$	6	5	4	3	2
1	2.12×10-77	2-04 × 16-19	2.04×10-18J	1.94710-19)	1.63×10-1
	94.100	95.0nm	92.3000	ma 601	12200
2	4 . 745 × 10-19 J	UP 01 x \$77. p	UP1-01×10.P	C P1-01× 80.0	
	ullam	43400	43600	LSLenn	
3	1. 815×10-19	1.548×10.19)	C***01×00.1		
	1095000	12×20m	1175nm		
4	1.5 x x 10-20 J	4.88×10-20			
	2 433 nm	4073000			
5	2.67×10-201				
	2445				

3. Assignment of Lines in the Hydrogen Spectrum

Atomic Spectrum Of Hydrogen Lab Answers

Jerry D. Wilson

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