



Activity B: Classifying reactions

Get the Gizmo ready:

- Turn off Show summary and Show histograms.



Introduction: Chemical equations show how **compounds** and **elements** react with one another. An element is a substance consisting of one kind of atom, such as aluminum (Al) or oxygen gas (O_2). A compound is a substance made of more than one kind of atom, such as water (H_2O) or table salt (NaCl).

Question: How are chemical reactions classified?

1. **Match:** Most chemical reactions can be classified as one of four types. Using the chemical equations in the Gizmo as a guide, match the following definitions to the type of reaction.

| | |
|---|---|
| B | One reactant is broken down into two or more products. |
| E | A fuel combines with oxygen to produce carbon dioxide (CO_2) and water (H_2O). |
| A | Two or more reactants combine to form one product. |
| D | Two compounds react to form two different compounds. |
| C | A compound reacts with an element to form a new compound and a different element. |

| |
|------------------------------|
| A. Synthesis |
| B. Decomposition |
| C. Single replacement |
| D. Double replacement |
| E. Combustion |

2. **Practice:** Balance each of the chemical equations below. (Some equations may already be in balance.) In the space to the right, classify the reaction as a *synthesis*, *decomposition*, *single replacement*, *double replacement*, or *combustion*.

| | | | | | | | | | |
|----|--------------------------------|------------------------------|--------------------------------|----------------------------|--------------------------------|-------------------------|--------------------------------|----------------|--------------------|
| A. | <input type="text" value="1"/> | $\text{AgNO}_3 +$ | <input type="text" value="1"/> | $\text{KCl} \rightarrow$ | <input type="text" value="1"/> | $\text{AgCl} +$ | <input type="text" value="1"/> | KNO_3 | Two replacements |
| B. | <input type="text" value="1"/> | $\text{H}_2\text{O} +$ | <input type="text" value="1"/> | $\text{SO}_3 \rightarrow$ | <input type="text" value="1"/> | H_2SO_4 | | | Synthesis |
| C. | <input type="text" value="2"/> | $\text{KI} +$ | <input type="text" value="1"/> | $\text{Cl}_2 \rightarrow$ | <input type="text" value="2"/> | $\text{KCl} +$ | <input type="text" value="1"/> | I_2 | Single replacement |
| D. | <input type="text" value="2"/> | $\text{NaHCO}_3 \rightarrow$ | <input type="text" value="1"/> | $\text{Na}_2\text{CO}_3 +$ | <input type="text" value="1"/> | $\text{H}_2\text{O} +$ | <input type="text" value="1"/> | CO_2 | Decomposition |
| E. | <input type="text" value="1"/> | $\text{Zn} +$ | <input type="text" value="2"/> | $\text{HCl} \rightarrow$ | <input type="text" value="1"/> | $\text{ZnCl}_2 +$ | <input type="text" value="1"/> | H_2 | Single replacement |

Balancing Chemical Equations Gizmo Answer Key

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