D.	2	KCIO <sub>3</sub> 🗆	2	KCI+	3	O <sub>2</sub>		
E.	2	Al +	3	CuO 🗆	1	Al <sub>2</sub> O <sub>3</sub> +	3	Cu
· F.	1	I <sub>2</sub> +	2	Na₂S₂O₃ □	2	Nal +	1	Na <sub>2</sub> S <sub>4</sub> O <sub>6</sub>
G.	6	Mg +	4	P. 0	2	Mg <sub>3</sub> P <sub>2</sub>		

Activity B:	Get the Gizmo ready:	
Classifying reactions	Turn off Show summary and Show histograms.	<i>₽</i>

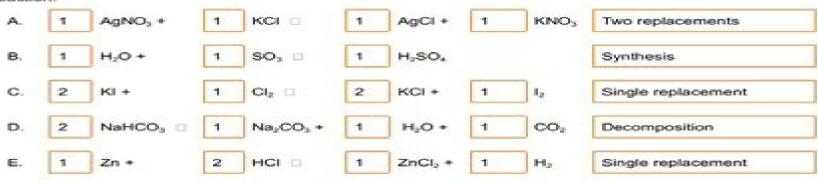
Introduction: Chemical equations show how compounds and elements react with one another. An element is a substance consisting of one kind of atom, such as aluminum (Al) or oxygen gas (O<sub>2</sub>). A compound is a substance made of more than one kind of atom, such as water (H<sub>2</sub>O) or table salt (NaCl).

#### Question: How are chemical reactions classified?

 Match: Most chemical reactions can be classified as one of four types. Using the chemical equations in the Gizmo as a guide, match the following definitions to the type of reaction.

В	One reactant is broken down into two or more products.	A. Synthesis
E	A fuel combines with oxygen to produce carbon dioxide (CO <sub>2</sub> ) and water (H <sub>2</sub> O).	B. Decomposition
,Α,	Two or more reactants combine to form one product.	C. Single replacement
D.	Two compounds react to form two different compounds.	D. Double replacement
C	A compound reacts with an element to form a new compound and a different element.	E. Combustion

 Practice: Balance each of the chemical equations below. (Some equations may already be in balance.) In the space to the right, classify the reaction as a synthesis, decomposition, single replacement, double replacement, or combustion.



# **Balancing Chemical Equations Gizmo Answer Key**

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