

Once this is done, the mass of the gas can be computed from the amount and the molar mass:

$$m = (1.66 \times 10^2 \text{ mol}) \times (16.04 \text{ g mol}^{-1}) = 2.67 \times 10^3 \text{ g} = \boxed{2.67 \times 10^3 \text{ kg}}$$

E1.5(b) Use the equation for hydrostatic pressure [1.3], $p = p_{\text{ex}} + \rho gh$. Let p_{ex} be the pressure at the top of the straw and p is the pressure on the surface of the liquid (atmospheric pressure). Thus, the pressure difference is

$$\begin{aligned} p - p_{\text{ex}} &= \rho gh = (1.0 \text{ g cm}^{-3}) \times \frac{1 \text{ kg}}{10^3 \text{ g}} \times \left(\frac{1 \text{ cm}}{10^{-2} \text{ m}} \right)^3 \times (9.81 \text{ m s}^{-2}) \times (0.15 \text{ m}) \\ &= \boxed{1.5 \times 10^3 \text{ Pa}} = 1.5 \times 10^{-2} \text{ atm} \end{aligned}$$

E1.6(b) The pressure in the apparatus is given by

$$p = p_{\text{ex}} + \rho gh \quad [1.3]$$

where $p_{\text{ex}} = 760 \text{ Torr} = 1 \text{ atm} = 1.013 \times 10^5 \text{ Pa}$,

and $\rho gh = 13.55 \text{ g cm}^{-3} \times \left(\frac{1 \text{ kg}}{10^3 \text{ g}} \right) \times \left(\frac{1 \text{ cm}}{10^{-2} \text{ m}} \right)^3 \times 0.100 \text{ m} \times 9.806 \text{ m s}^{-2} = 1.33 \times 10^4 \text{ Pa}$

$$p = 1.013 \times 10^5 \text{ Pa} + 1.33 \times 10^4 \text{ Pa} = 1.146 \times 10^5 \text{ Pa} = \boxed{115 \text{ kPa}}$$

E1.7(b) Rearrange the perfect gas equation [1.8] to give $R = \frac{pV}{nT} = \frac{pV_{\text{m}}}{T}$

All gases are perfect in the limit of zero pressure. Therefore, the value of pV_{m}/T extrapolated to zero pressure will give the best value of R .

The molar mass can be introduced through

$$pV = nRT = \frac{m}{M} RT$$

which on rearrangement gives $M = \frac{m}{V} \frac{RT}{p} = \rho \frac{RT}{p}$.

The best value of M is obtained from an extrapolation of ρ/p versus p to zero pressure; the intercept is M/RT .

Draw up the following table:

p/atm	$(pV_{\text{m}}/T)/(\text{dm}^3 \text{ atm K}^{-1} \text{ mol}^{-1})$	$(\rho/p)/(\text{g dm}^{-3} \text{ atm}^{-1})$
0.750 000	0.082 0014	1.428 59
0.500 000	0.082 0227	1.428 22
0.250 000	0.082 0414	1.427 90

From Figure 1.1(a), $\left(\frac{pV_{\text{m}}}{T} \right)_{p=0} = \boxed{0.082 062 \text{ dm}^3 \text{ atm K}^{-1} \text{ mol}^{-1}}$

Physical Chemistry 9th Edition Solutions

C. A. Trapp, M. P. Cady, Carmen Giunta



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Instructor's Solutions Manual to Accompany Atkins' Physical Chemistry, Ninth Edition C. A. Trapp, M. P.

Cady, Carmen Giunta, 2010 The Instructor's solutions manual to accompany Atkins Physical Chemistry provides detailed solutions to the odd exercises and the even numbered discussion questions and problems that feature in the ninth edition of Atkins Physical Chemistry The manual is intended for instructors and consists of material that is not available to undergraduates The manual is free to all adopters of the main text [Physico-chemical Aspects of Textile Coloration](#)

Stephen M. Burkinshaw, 2016-02-08 The production of textile materials comprises a very large and complex global industry that utilises a diverse range of fibre types and creates a variety of textile products As the great majority of such products are coloured predominantly using aqueous dyeing processes the coloration of textiles is a large scale global business in which complex procedures are used to apply different types of dye to the various types of textile material The development of such dyeing processes is the result of substantial research activity undertaken over many decades into the physico chemical aspects of dye adsorption and the establishment of dyeing theory which seeks to describe the mechanism by which dyes interact with textile fibres Physico Chemical Aspects of Textile Coloration provides a comprehensive treatment of the physical chemistry involved in the dyeing of the major types of natural man made and synthetic fibres with the principal types of dye The book covers fundamental aspects of the physical and chemical structure of both fibres and dyes together with the structure and properties of water in relation to dyeing dyeing as an area of study as well as the terminology employed in dyeing technology and science contemporary views of intermolecular forces and the nature of the interactions that can occur between dyes and fibres at a molecular level fundamental principles involved in dyeing theory as represented by the thermodynamics and kinetics of dye sorption detailed accounts of the mechanism of dyeing that applies to cotton and other cellulosic fibres polyester polyamide wool polyacrylonitrile and silk fibres non aqueous dyeing as represented by the use of air organic solvents and supercritical CO₂ fluid as alternatives to water as application medium The up to date text is supported by a large number of tables figures and illustrations as well as footnotes and widespread use of references to published work The book is essential reading for students teachers researchers and professionals involved in textile coloration

Thermodynamics of Pharmaceutical Systems Kenneth A. Connors, 2003-03-19 Studies of thermodynamics often fail to demonstrate how the mathematical intricacies of the subject relate to practical laboratory applications Thermodynamics of Pharmaceutical Systems makes these connections clear emphasizing specific applications to pharmaceutical systems in a study created specifically for contemporary curriculums at colleges of pharmacy Students investigating drug discovery drug delivery and drug action will benefit from Kenneth Connors's authoritative treatment of the fundamentals of thermodynamics as well as his attention to drug molecules and experimental considerations An extensive appendix that reviews the mathematics needed to master the pharmacy curriculum proves an invaluable reference Connors

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Solid State Chemistry Aaron Wold,Kirby Dwight,2012-12-06 The subject matter of solid state chemistry lies within the spheres of both physical and inorganic chemistry In addition there is a large overlap with solid state physics and materials engineering However solid state chemistry has still to be recognized by the general body of chemists as a legitimate subfield of chemistry The discipline is not even well defined as to content and has many facets that make writing a textbook a formidable task The early studies carried out in the United States by Roland Ward and his co workers emphasized the synthesis of new materials and the determination of their structure His work on doped alkaline earth sulfides formed the basis for the development of infrared phosphors and his pioneering studies on oxides were important in understanding the structural features of both the perovskite oxides as well as the magnetoplumbites In 1945 A F Wells published the first edition of *Structural Inorganic Chemistry* This work attempts to demonstrate that the synthesis structure and properties of solids form an important part of inorganic chemistry Now after almost 50 years during which many notable advances have been made in solid state chemistry it is still evident that the synthesis structure determination and properties of solids receive little attention in most treatments of inorganic chemistry The development of the field since the early studies of Roland Ward early 1940s has been rapid **Physical Chemistry An Advanced Treatise** Wilhelm Jost,2012-12-02 Physical

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Table of Contents Physical Chemistry 9th Edition Solutions

1. Understanding the eBook Physical Chemistry 9th Edition Solutions
 - The Rise of Digital Reading Physical Chemistry 9th Edition Solutions
 - Advantages of eBooks Over Traditional Books
2. Identifying Physical Chemistry 9th Edition Solutions
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Physical Chemistry 9th Edition Solutions
 - User-Friendly Interface
4. Exploring eBook Recommendations from Physical Chemistry 9th Edition Solutions
 - Personalized Recommendations
 - Physical Chemistry 9th Edition Solutions User Reviews and Ratings

- Physical Chemistry 9th Edition Solutions and Bestseller Lists
- 5. Accessing Physical Chemistry 9th Edition Solutions Free and Paid eBooks
 - Physical Chemistry 9th Edition Solutions Public Domain eBooks
 - Physical Chemistry 9th Edition Solutions eBook Subscription Services
 - Physical Chemistry 9th Edition Solutions Budget-Friendly Options
- 6. Navigating Physical Chemistry 9th Edition Solutions eBook Formats
 - ePub, PDF, MOBI, and More
 - Physical Chemistry 9th Edition Solutions Compatibility with Devices
 - Physical Chemistry 9th Edition Solutions Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Physical Chemistry 9th Edition Solutions
 - Highlighting and Note-Taking Physical Chemistry 9th Edition Solutions
 - Interactive Elements Physical Chemistry 9th Edition Solutions
- 8. Staying Engaged with Physical Chemistry 9th Edition Solutions
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Physical Chemistry 9th Edition Solutions
- 9. Balancing eBooks and Physical Books Physical Chemistry 9th Edition Solutions
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Physical Chemistry 9th Edition Solutions
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Physical Chemistry 9th Edition Solutions
 - Setting Reading Goals Physical Chemistry 9th Edition Solutions
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Physical Chemistry 9th Edition Solutions
 - Fact-Checking eBook Content of Physical Chemistry 9th Edition Solutions
 - Distinguishing Credible Sources

13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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