Once this is done, the mass of the gas can be computed from the amount and the molar mass:

$$m = (1.66 \times 10^{5} \text{ mol}) \times (16.04 \text{ g mol}^{-1}) = 2.67 \times 10^{6} \text{ g} = 2.67 \times 10^{3} \text{ kg}$$

Use the equation for hydrostatic pressure [1.3], $p = p_m + pgh$. Let p_m be the pressure at the top of the straw and p is the pressure on the surface of the liquid (atmospheric pressure). Thus, the pressure difference is

$$p - p_{\text{ex}} = \rho g h = (1.0 \text{ g cm}^{-3}) \times \frac{1 \text{ kg}}{10^3 \text{ g}} \times \left(\frac{1 \text{ cm}}{10^{-2} \text{ m}}\right)^3 \times (9.81 \text{ m s}^{-2}) \times (0.15 \text{ m})$$
$$= \left[1.5 \times 10^3 \text{ Pa}\right] = 1.5 \times 10^{-2} \text{ atm}$$

E1.6(b) The pressure in the apparatus is given by

$$p = p_{as} + \rho g h [1.3]$$

where $p_{ex} = 760 \text{ Torr} = 1 \text{ atm} = 1.013 \times 10^5 \text{ Pa}$,

and
$$\rho g h = 13.55 \text{ g cm}^{-3} \times \left(\frac{1 \text{ kg}}{10^3 \text{ g}}\right) \times \left(\frac{1 \text{ cm}}{10^{-2} \text{ m}}\right)^3 \times 0.100 \text{ m} \times 9.806 \text{ m s}^{-2} = 1.33 \times 10^4 \text{ Pa}$$

$$\rho = 1.013 \times 10^5 \text{ Pa} + 1.33 \times 10^4 \text{ Pa} = 1.146 \times 10^5 \text{ Pa} = \boxed{115 \text{ kPa}}$$

E1.7(b) Rearrange the perfect gas equation [1.8] to give $R = \frac{pV}{nT} = \frac{pV_{m}}{T}$

All gases are perfect in the limit of zero pressure. Therefore, the value of pV_m/T extrapolated to zero pressure will give the best value of R.

The molar mass can be introduced through

$$pV = nRT = \frac{m}{M}RT$$

which on rearrangement gives $M = \frac{m}{V} \frac{RT}{\rho} = \rho \frac{RT}{\rho}$.

The best value of M is obtained from an extrapolation of ρ/ρ versus ρ to zero pressure; the intercept is M/RT.

Draw up the following table:

p/atm	$(pV_{\infty}/T)/(dm^{5} atm K^{-1} mol^{-1})$	$(\rho/p)/(g \text{ dm}^{-3} \text{ atm}^{-3})$
0.750 000	0.082 0014	1.428 59
0.500 000	0.082 0227	1,428 22
0.250 000	0.082 0414	1.427 90

From Figure 1.1(a),
$$\left(\frac{pV_m}{T}\right)_{p=0} = \left[0.082\ 062\ dm^3\ atm\ K^{-1}\ mol^{-1}\right]$$

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