$$M = RT \left(\frac{\rho}{\rho}\right)_{\rho=0} = (0.082\,061\,5\,\text{L.atm}\,\text{mol}^{-1}\,\text{K}^{-1}) \times (273.15\,\text{K}) \times (1.42755\,\text{g.L}^{-1}\,\text{atm}^{-1})$$
  
=  $31.9987\,\text{g.mol}^{-1}$ 

The value obtained for R deviates from the accepted value by 0.005 per cent. The error results from the fact that only three data points are available and that a linear extrapolation was employed. The molar mass, however, agrees exactly with the accepted value, probably because of compensating plotting errors.

E1.11(b) The mass density  $\rho$  is related to the molar volume  $V_m$  by

$$V_m = \frac{M}{\rho}$$

where M is the molar mass. Putting this relation into the perfect gas law yields

$$pV_m = RT$$
 so  $\frac{pM}{\rho} = RT$ 

Rearranging this result gives an expression for M; once we know the molar mass, we can divide by the molar mass of phosphorus atoms to determine the number of atoms per gas molecule

$$M = \frac{RT\rho}{\rho} = \frac{(62.364\,\mathrm{L\,Torr\,K^{-1}\,mol^{-1}}) \times [(100 + 273)\,\mathrm{K}] \times (0.6388\,\mathrm{g\,L^{-1}})}{120\,\mathrm{Torr}} = 124\,\mathrm{g\,mol^{-1}}.$$

The number of atoms per molecule is

$$\frac{124 \text{ g mol}^{-1}}{31.0 \text{ g mol}^{-1}} = 4.00$$

suggesting a formula of P4

E1.12(b) Use the perfect gas equation to compute the amount; then convert to mass.

$$pV = nRT$$
 so  $n = \frac{pV}{RT}$ 

We need the partial pressure of water, which is 53 per cent of the equilibrium vapour pressure at the given temperature and standard pressure.

$$p = (0.53) \times (2.69 \times 10^{3} \text{ Pa}) = 1.43 \times 10^{3} \text{ Pa}$$
  
so  $n = \frac{(1.43 \times 10^{3} \text{ Pa}) \times (250 \text{ m}^{3})}{(8.3145 \text{ J K}^{-1} \text{ mol}^{-1}) \times (23 + 273) \text{ K}} = 1.43 \times 10^{2} \text{ mol}$   
or  $m = (1.43 \times 10^{2} \text{ mol}) \times (18.0 \text{ g mol}^{-1}) = 2.61 \times 10^{3} \text{ g} = 2.61 \text{ kg}$ 

E1.13(b) (a) The volume occupied by each gas is the same, since each completely fills the container. Thus solving for V from eqn 14 we have (assuming a perfect gas)

$$\begin{split} V &= \frac{n_I R T}{p_I} - n_{\rm Ne} = \frac{0.225 \ \mathrm{g}}{20.18 \ \mathrm{g \ mol}^{-1}} \\ &= 1.11\overline{5} \times 10^{-2} \ \mathrm{mol}, \quad p_{\rm Ne} = 66.5 \ \mathrm{Torr}, \qquad T = 300 \ \mathrm{K} \\ V &= \frac{(1.11\overline{5} \times 10^{-2} \ \mathrm{mol}) \times (62.36 \ \mathrm{L \ Torr \ K}^{-1} \ \mathrm{mol}^{-1}) \times (300 \ \mathrm{K})}{66.5 \ \mathrm{Torr}} = 3.137 \ \mathrm{L} = \boxed{3.14 \ \mathrm{L}} \end{split}$$

# **Physical Chemistry Atkins 7th Edition Solutions Manual**

**JE Gale** 

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