

Study Buddy: Atomic Structure and Nuclear Chemistry

Development of the Modern Atomic Theory

Scientist	Experiment	Findings	Limitations	Model
Democritus	No experiment	Atoms are the smallest form of matter and cannot be created, destroyed or divided	Did not explain chemical behavior	
Dalton	Use experiments of previous scientists	Postulates: Atoms of elements are identical; atoms can't be created or destroyed; atoms combine in whole number ratios in a chemical reaction	Did not include sub-atomic particles	
Thomson	Cathode ray	Discovered the electron "Plum-pudding" Model	Did not explain deflection pattern of alpha particles as seen by Rutherford	
Rutherford	Gold foil	Discovered the proton and nucleus. Atoms contain mostly empty space.	Could not explain why electrons do not crash into the nucleus	
Chadwick	Alpha particles	Discovered the neutron	No modification of atom model	No new model/neutron later add to nucleus
Bohr	Uses findings from previous scientists and Quantum theory	Electrons are in orbits of differing energy levels.	Does not explain the behavior of other elements just hydrogen	
Modern Atom	Multiple experiments	Electron cloud and dual nature of electrons	Still doesn't explain the exceptions.	

Parts of an Atom (8th Grade Review)

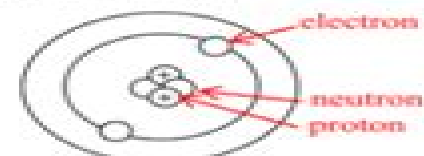
Sub-atomic Particle	Location	Charge	Mass
Proton (p ⁺)	Nucleus	1+	1
Neutron (n ⁰)	Nucleus	0	1
Electron (e ⁻)	outside the nucleus	1-	0 or 1/1840

Nuclear Symbol Carbon-14



Protons:	6
Neutrons:	8
Atomic Number:	6
Mass Number:	14

Identify the proton, electron, neutron in the model below.



What is the atomic number of the above models? 2
What is the atomic mass? 4.00 amu
What element is this? Helium

Atomic Structure Study Guide And Nuclear Processes

Marco Cascella

A decorative graphic consisting of a red circular shape with a white center, partially obscured by a white horizontal bar that extends from the left edge of the slide.

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