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# FOCUS 2 MOLECULES

- 2A.2 (a) 5; (b) 8; (c) 12; (d) 7
- 2A.4 (a) [Ar] 3d<sup>1</sup>; (b) [Ar] 3d<sup>5</sup>; (c) [Kr] 4d<sup>10</sup>; (d) [Xe]
- 2A.6 (a) [Kr]; (b) [Xe] 4f<sup>14</sup>5d<sup>5</sup>; (c) [Xe]; (d) [Ar]
- 2A.8 (a) Ca: [Ar] 4s², Ti²\*: [Ar] 3d²; V³\*: [Ar] 3d² In the d block, the energies of the n-1 d-orbitals lie below those of the ns-orbitals. Therefore, when V and Ti form ions they lose their 3s electrons before losing their 4d electrons. (b) Ca: no unpaired electrons; Ti²\*: two unpaired electrons; V³\*: two unpaired electrons. (c) Ti³\*: [Ar] 3d¹ no neutral atom has this electron configuration.
- 2A.10 (a) Au3+; (b) Os3+; (c) 13+; (d) As3+
- 2A.12 (a) Cr2+; (b) Ag2+; (c) Zn2+; (d) Pb2+
- 2A.14 (a) 4d; (b) 3p; (c) 5p; (d) 4s
- **2A.16** (a) +1; (b) -2; (c) +2; (d) -3; (e) -1

# **Atkins 7th Edition Solutions Manual**

C. A. Trapp, M. P. Cady, Carmen Giunta



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