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**Jeff Killingsworth, John Harold Haynes**

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web Jan 10 2023 introduction in this lab we will apply the concepts of reaction rate order of reaction and catalyst our example reaction is the iodination of acetone catalyzed by hydrochloric acid the reaction is zero order in iodine you will figure out the order in acetone and the order in HCl and the rate constant

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web video answer there can be a 6.0 M solution in sodium hydroxide ensure it has been reduced to 400 liters so we have to find the moles in the solution the first thing

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web the iodination of acetone is convenient because the color of the iodine reactant is readily observable and the change in concentration of the iodine can be easily monitored by our spectrophotometer and the reaction proceeds quickly enough for multiple trials but slowly enough to obtain measurements

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web in the iodination of acetone the rate of this reaction is found to depend on the concentration of hydrogen ion in the



solution as well as pre sumably on the concentrations of the two reactants

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web video answer the data was given by an point we arrange the dates in ascending order this is the pointing we need to build a stem and plot graph to present quantitiy data included a graphical input this is the stem part and this is the leaf item yours is

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