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FOCUS 2

MOLECULES

2A.2 (a) 5; (b) 8; (c) 12; (d) 7

2A.4 (a) $[\text{Ar}] 3d^1$; (b) $[\text{Ar}] 3d^5$; (c) $[\text{Kr}] 4d^{10}$; (d) $[\text{Xe}]$

2A.6 (a) $[\text{Kr}]$; (b) $[\text{Xe}] 4f^{14}5d^5$; (c) $[\text{Xe}]$; (d) $[\text{Ar}]$

2A.8 (a) Ca: $[\text{Ar}] 4s^2$, Ti^{2+} : $[\text{Ar}] 3d^2$; V^{3+} : $[\text{Ar}] 3d^2$ In the d block, the energies of the $n-1$ d-orbitals lie below those of the ns -orbitals. Therefore, when V and Ti form ions they lose their 3s electrons before losing their 4d electrons. (b) Ca: no unpaired electrons; Ti^{2+} : two unpaired electrons; V^{3+} : two unpaired electrons. (c) Ti^{3+} : $[\text{Ar}] 3d^1$ no neutral atom has this electron configuration.

2A.10 (a) Au^{3+} ; (b) Os^{3+} ; (c) I^{3+} ; (d) As^{3+}

2A.12 (a) Cr^{2+} ; (b) Ag^{2+} ; (c) Zn^{2+} ; (d) Pb^{2+}

2A.14 (a) 4d; (b) 3p; (c) 5p; (d) 4s

2A.16 (a) +1; (b) -2; (c) +2; (d) -3; (e) -1

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Peter Atkins



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The British National Bibliography Arthur James Wells, 2002

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