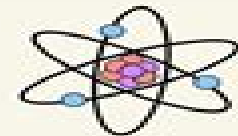


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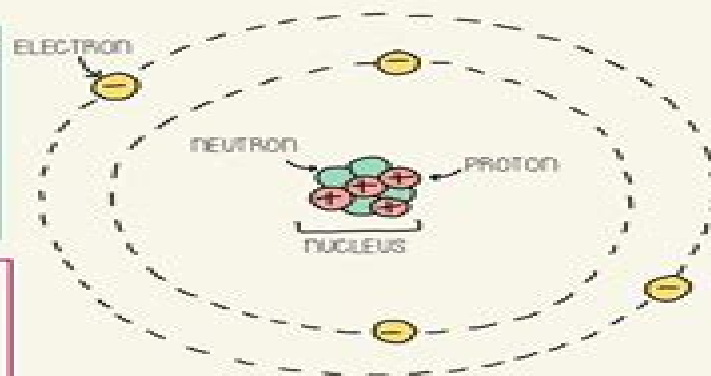
- Atomic structure -

Every atom is made up of a central **nucleus** surrounded by **electrons** whizzing around the nucleus in electron **orbits**. The nucleus is tiny in relation to the size of the whole atom - comparable to a pea in the middle of a football stadium. The nucleus is made up of **protons and neutrons**, which are the heaviest particles. Electrons are approximately 2000 times lighter than a proton or neutron and weigh practically nothing, therefore most of the mass comes from the tiny nucleus. **Protons** are **positively charged** whereas **neutrons** have **no charge**, making the nucleus positively charged overall. **Electrons** are **negatively charged**, which enables the nucleus to keep hold of its electrons, due to the attraction between the opposite charges.

- Atomic number & mass number -

Each atom has an **atomic number** and a **mass number**, which are displayed on the Periodic table as the small number and larger number next to each chemical symbol. The **atomic number** tells us the number of **protons** (which is unique for each element. If the number of protons changes (and therefore the atomic number changes), which could occur during radioactive decay, for example, then the identity of the element also changes).

The **mass number** tells us the number of both **protons and neutrons** in the nucleus. Remember that most of an atom's mass comes from the nucleus and that the nucleus is made up of both protons and neutrons. Two atoms of the same element can have different mass numbers because they are isotopes.



Atomic number →

Mass number →



Atomic particle	Charge	Mass
Proton	+1	1
Neutron	0	1
Electron	-1	1/2000

- Isotopes -

Isotopes are atoms which have the **same number of protons** but **different numbers of neutrons**. Therefore they have the same atomic number but a different mass number. For example, the most common isotopes of carbon are carbon-12 and carbon-13, in which carbon-13 has an additional neutron in its nucleus to give it a higher mass number.

The mass numbers that we see in the periodic table are an average of all the masses of all isotopes of a particular element. We call this the **relative atomic mass (RAM)** and can be calculated using the following formula:

$$\frac{(\text{Mass number of isotope 1} \times \text{percentage of isotope 1}) + (\text{mass number of isotope 2} \times \text{percentage of isotope 2})}{100}$$

Worked example

Let's say we have a sample of carbon which consists of 76% of the isotope carbon-12 and 24% carbon-13, then the relative atomic mass (RAM) would be:

$$\text{RAM} = \frac{(76 \times 12) + (24 \times 13)}{100}$$

$$\text{RAM} = 12.24$$

Questions

1. Name the two subatomic particles that are found in the nucleus and state their charges.
2. What does atomic number represent?
3. Why do atoms have an overall neutral charge?
4. Sodium (Na) has an atomic number of 11 and a mass number of 23. State the number of protons, neutrons and electrons in a sodium atom.
5. What is an isotope?
6. A sample of bromine contains 45% bromine-79 and 55% bromine-81. Calculate its relative atomic mass (RAM).

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