

STUDY GUIDE - Atomic Structure

History of Atom

- Gold Foil Experiment and discovery of nucleus and proton

Atomic Structure

- Know the subatomic particles, their location, their charge, their relative mass
- Be able to calculate the # of
 - Protons
 - Neutrons (using Mass # and/or Atomic Mass)
 - Electrons

Atomic Mass, Atomic Number and Mass Number

- Understand what the difference is between them
- Be able to calculate an atomic mass from the percent abundance and mass of the isotopes of that element

Isotope

- Be able to write both forms of Isotope Notation
- Explain what the difference is between isotopes

Ion

- What it means for an ion to have a positive charge
- What it means for an ion to have a negative charge
- Be able to calculate the charge and the number of electrons in an ion from the charge

Definitions

Atom

Proton

Neutron

Electron

Nucleus

Isotope

Mass number

Atomic mass

Energy level

Ion

Atomic Number

Atomic Structure Answer Key Study Guide

Bruce A. Chase, Peter J. Russell



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