

EQ: Describe atomic structure.

- Atoms are made of extremely tiny particles called protons, neutrons, and electrons.
- Protons and neutrons are in the center of the atom, making up the nucleus.
- Electrons surround the nucleus.
- Protons have a positive charge.
- Electrons have a negative charge.
- The charge on the proton and electron are exactly the same size but opposite.
- Neutrons have no charge.
- Since opposite charges attract protons and electrons attract each other.

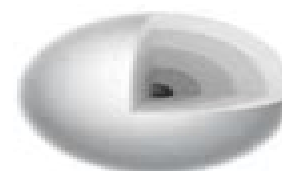
- ☐ Label the nucleus (protons, neutrons) and electrons in the drawing of a carbon atom.
- ☐ Draw a line between the subatomic particle and its charge.

proton
electron
neutron

no charge
positive charge
negative charge



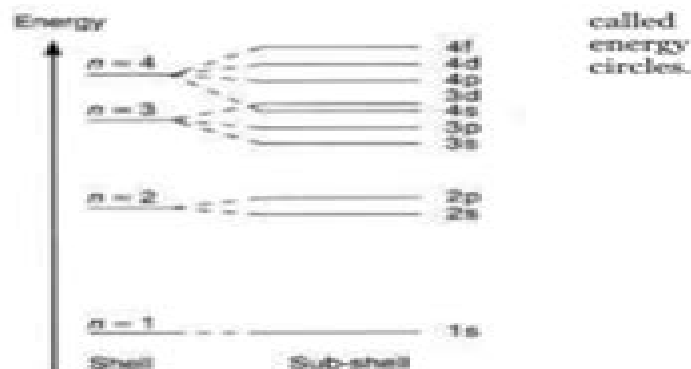
- The electrons surrounding an atom are located in regions around the nucleus called "energy levels".
- An energy level represents the 3-dimensional space surrounding the nucleus where electrons are most likely to be.
- The first energy level is closest to the nucleus. The second energy level is a little farther away than the first. The third is a little farther away than the second, and so on.
- Each energy level can accommodate or "hold" a different number of electrons before additional electrons begin to go into the next level.
- When the first energy level has 2 electrons, the next electrons go into the second energy level until the second level has 8 electrons.
- When the second energy level has 8 electrons, the next electrons go into the third energy level until the third level has 8 electrons.
- When the third energy level has 8 electrons, the next 2 electrons go into the fourth energy level.
- The electrons in the energy level farthest from the nucleus are called valence electrons.
- Atoms in the same column (group) in the periodic table have the same number of valence electrons.



Energy levels

Electrons surround the nucleus of an atom in regions *energy levels*. Even though atoms are spherical, the levels in an atom are more easily shown in concentric

Energy Level	Maximum # electrons
1	2
2	8
3	18
4	32
5	50
6	72



Atomic Structure Study Guide

David R. Klein



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This book delves into Atomic Structure Study Guide. Atomic Structure Study Guide is a crucial topic that needs to be grasped by everyone, ranging from students and scholars to the general public. The book will furnish comprehensive and in-depth insights into Atomic Structure Study Guide, encompassing both the fundamentals and more intricate discussions.

1. This book is structured into several chapters, namely:

- Chapter 1: Introduction to Atomic Structure Study Guide
- Chapter 2: Essential Elements of Atomic Structure Study Guide
- Chapter 3: Atomic Structure Study Guide in Everyday Life
- Chapter 4: Atomic Structure Study Guide in Specific Contexts
- Chapter 5: Conclusion

2. In chapter 1, this book will provide an overview of Atomic Structure Study Guide. The first chapter will explore what Atomic Structure Study Guide is, why Atomic Structure Study Guide is vital, and how to effectively learn about Atomic Structure Study Guide.
3. In chapter 2, this book will delve into the foundational concepts of Atomic Structure Study Guide. The second chapter will elucidate the essential principles that need to be understood to grasp Atomic Structure Study Guide in its entirety.
4. In chapter 3, the author will examine the practical applications of Atomic Structure Study Guide in daily life. This chapter will showcase real-world examples of how Atomic Structure Study Guide can be effectively utilized in everyday scenarios.
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6. In chapter 5, the author will draw a conclusion about Atomic Structure Study Guide. The final chapter will summarize the key points that have been discussed throughout the book.

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