Name:

 Hydrogen, H₂, is used as rocket fuel. The hydrogen is burned in oxygen, O₂, according to the following equation. What is the enthalpy change if 200.0 g of hydrogen, H2 is burned?

2 $H_2(g) + O_2(g) \rightarrow 2 H_2O(g)$; $\Delta H = -484 \text{ kJ}$

200.03 12 1 mol HZ = 99.2 mol Hz 99.2 mol Hz 1-484 KJ (2011)

Carbon disulfide burns in air, producing carbon dioxide and sulfur dioxide

 $CS_2(l) + 3O_2(g) \rightarrow CO_2(g) + 2SO_2(g) : \Delta H = -1077 \text{ kJ}$ What is ΔH for the following reaction?

change 50 5:5 Not 1/2 CO2(8) + SO2(8) - 1/2 CS2(1) + 3/2 O2(8)

+ 1077 K3 = (538.5 K3)

3. Following is the equation for the combustion of acetone:

2 C₃H₆O(g) + 8 O₂(g) - 6 H₂O(g) + 6 CO₂(g) + 853.6 kcal/2

How much heat will be given off (in kcal) if 4.5 mol of acetone is burned completely?

- 853. 6 Kc-1 2 mol Catte a

4.5 mol (3400 = (-1920.6 KC-1)

 The specific heat of iron is 0.110 cal / g-°C. How much heat must be added to 5.2 g of iron at 25 °C to raise its temperature to 225 °C? heat = 8 = m e AT = (5.2) (0.110 cml) (225-25'C) = (14.4)

 Hydrazine, N2H4, is a colorless liquid used as a rocket fuel. What is the enthalpy change for the Grant process in which hydrazine is formed from its elements (reaction (1) below)? Use Hess's Law to calculate ΔH for reaction (1) given the data below.

> (1) $N_2(g) + 2H_2(g) \rightarrow$

 $N_2H_4(1) + O_2(g) \rightarrow N_2(g) + 2H_2O(1); \Delta H = -622.2 \text{ kJ}$

 $H_2(g) + \frac{1}{2}O_2(g) \rightarrow H_2O(f) : \Delta H = -285.8 \text{ kJ}$

N291+ 2420 (1) --> N2Hy (1) + O2(1, DH=622.2)

2H2G1 + OZG1 -> ZHZOQ1 DH= (Z)(-285.8) KJ NZG1 + ZHZG1 -> NZHYG1 (DH= + 50.6 KJ

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