#### BONDING REVIEW

1. An atom has 6 total electrons. How many electrons will be involved in bonding with another atom?

2. Intramolecular bonds are bonds \_\_\_\_\_\_ molecules while Intermolecular bonds are bonds

D where

What is Electronegativity?

4. Fill in this table for Intramolecular bonds

TYPE OF BOND	Eneg Diff	Charges produced	2 EXAMPLES	
Ionic	1.7 or greate	r Strong	Noce -any pie	of ata
Polar covalent	05-67	slight golf-	H <sub>2</sub> O, NH <sub>3</sub>	
Non-polar covalent	Ges than		or any	

5. For a molecule to be considered polar it needs to have:

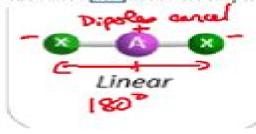
· polar bondo (Janic or polar combat)

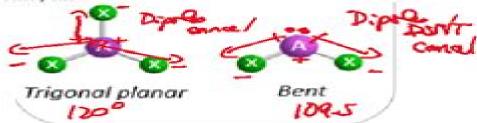
amd

" Dipoles that Dist and out

6. Look at the Molecular geometries below.

Oircle the VSEPR shapes that could cause the dipoles to cancel out. <u>Underline</u> the shapes with a bond angle of 120°. Put a box around the shapes that would have lone pairs.





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**L** Manion



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