

BONDING REVIEW

1. An atom has 6 total electrons. How many electrons will be involved in bonding with another atom?
2. Intramolecular bonds are bonds within molecules while Intermolecular bonds are bonds between molecules.
3. What is Electronegativity?

④ only valence e⁻ are involved in bonds

4. Fill in this table for Intramolecular bonds

| TYPE OF BOND | Eneg Diff | Type of charges produced | 2 EXAMPLES |
|--------------------|----------------|------------------------------|-----------------------------------------------|
| Ionic | 1.7 or greater | Strong | NaCl - any pair of atoms with |
| Polar covalent | 0.5 - 1.7 | slight δ^+ δ^- | KCl H ₂ O, NH ₃ |
| Non-polar covalent | less than 0.5 | none | O ₂ , Cl ₂ , etc or any |

5. For a molecule to be considered polar it needs to have:

- i. polar bonds (Ionic or polar covalent)

and

- ii. Dipoles that DON'T cancel out

6. Look at the Molecular geometries below.

Circle the VSEPR shapes that could cause the dipoles to cancel out. Underline the shapes with a bond angle of 120°. Put a box around the shapes that would have lone pairs.



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