

BONDING REVIEW

1. An atom has 6 total electrons. How many electrons will be involved in bonding with another atom?
2. Intramolecular bonds are bonds within molecules while Intermolecular bonds are bonds between molecules.
3. What is Electronegativity?

④ only valence e⁻ are involved in bonds

4. Fill in this table for Intramolecular bonds

TYPE OF BOND	Eneg Diff	Type of charges produced	2 EXAMPLES
Ionic	1.7 or greater	Strong	NaCl - any pair of atoms with
Polar covalent	0.5 - 1.7	slight δ^+ δ^-	KCl H ₂ O, NH ₃
Non-polar covalent	less than 0.5	none	O ₂ , Cl ₂ , etc or any

5. For a molecule to be considered polar it needs to have:

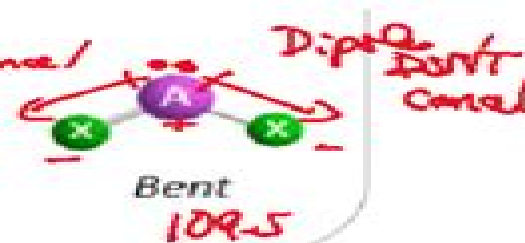
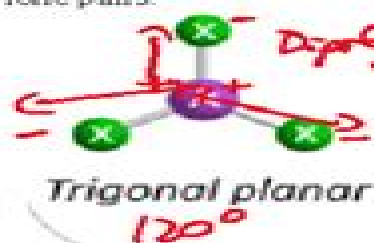
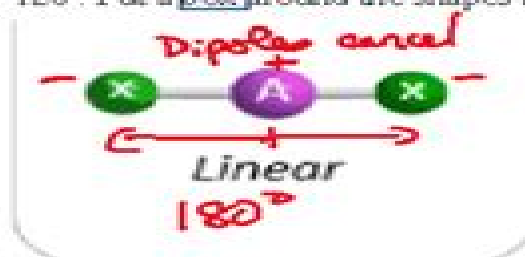
- i. polar bonds (Ionic or polar covalent)

and

- ii. Dipoles that DON'T cancel out

6. Look at the Molecular geometries below.

Circle the VSEPR shapes that could cause the dipoles to cancel out. Underline the shapes with a bond angle of 120°. Put a box around the shapes that would have lone pairs.



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Table of Contents Atom And Bonding Review And Reinforce

1. Understanding the eBook Atom And Bonding Review And Reinforce
 - The Rise of Digital Reading Atom And Bonding Review And Reinforce
 - Advantages of eBooks Over Traditional Books
2. Identifying Atom And Bonding Review And Reinforce
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Atom And Bonding Review And Reinforce
 - User-Friendly Interface
4. Exploring eBook Recommendations from Atom And Bonding Review And Reinforce
 - Personalized Recommendations
 - Atom And Bonding Review And Reinforce User Reviews and Ratings
 - Atom And Bonding Review And Reinforce and Bestseller Lists
5. Accessing Atom And Bonding Review And Reinforce Free and Paid eBooks

- Atom And Bonding Review And Reinforce Public Domain eBooks
- Atom And Bonding Review And Reinforce eBook Subscription Services
- Atom And Bonding Review And Reinforce Budget-Friendly Options
- 6. Navigating Atom And Bonding Review And Reinforce eBook Formats
 - ePub, PDF, MOBI, and More
 - Atom And Bonding Review And Reinforce Compatibility with Devices
 - Atom And Bonding Review And Reinforce Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Atom And Bonding Review And Reinforce
 - Highlighting and Note-Taking Atom And Bonding Review And Reinforce
 - Interactive Elements Atom And Bonding Review And Reinforce
- 8. Staying Engaged with Atom And Bonding Review And Reinforce
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Atom And Bonding Review And Reinforce
- 9. Balancing eBooks and Physical Books Atom And Bonding Review And Reinforce
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Atom And Bonding Review And Reinforce
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Atom And Bonding Review And Reinforce
 - Setting Reading Goals Atom And Bonding Review And Reinforce
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Atom And Bonding Review And Reinforce
 - Fact-Checking eBook Content of Atom And Bonding Review And Reinforce
 - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development

- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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