

Data:

Experiment 6: The Atomic Spectrum of Hydrogen

1. Calculate the Values of the Energy Levels of the Hydrogen Atom

n	Energy, E_n	n	Energy, E_n
1	$-2.18 \times 10^{-18} \text{ J}$	4	$-1.36 \times 10^{-19} \text{ J}$
2	$-5.45 \times 10^{-19} \text{ J}$	5	$-8.72 \times 10^{-20} \text{ J}$
3	$-2.42 \times 10^{-19} \text{ J}$	6	$-6.05 \times 10^{-20} \text{ J}$

2. Calculation of the Wavelengths of the Lines in the Hydrogen Spectrum

In the upper half of each box write the difference in energy between $E_{n_{\text{hi}}}$ and $E_{n_{\text{lo}}}$. In the lower half of the box write the wavelength associated with that value.

$n_{\text{hi}} \downarrow$ $n_{\text{lo}} \rightarrow$	6	5	4	3	2
1	$2.12 \times 10^{-17} \text{ J}$ 94.6 nm	$2.09 \times 10^{-17} \text{ J}$ 95.0 nm	$2.04 \times 10^{-18} \text{ J}$ 97.3 nm	$1.94 \times 10^{-19} \text{ J}$ 103 nm	$1.63 \times 10^{-19} \text{ J}$ 122 nm
2	$4.845 \times 10^{-19} \text{ J}$ 411 nm	$4.578 \times 10^{-19} \text{ J}$ 434 nm	$4.09 \times 10^{-19} \text{ J}$ 486 nm	$3.03 \times 10^{-20} \text{ J}$ 656 nm	
3	$1.815 \times 10^{-19} \text{ J}$ 1095 nm	$1.548 \times 10^{-19} \text{ J}$ 1282 nm	$1.06 \times 10^{-19} \text{ J}$ 1875 nm		
4	$7.55 \times 10^{-20} \text{ J}$ 2633 nm	$4.88 \times 10^{-20} \text{ J}$ 4073 nm			
5	$2.67 \times 10^{-20} \text{ J}$ 7445 nm				

3. Assignment of Lines in the Hydrogen Spectrum

Atomic Spectrum Of Hydrogen Pre Lab Answers

Scott C. Dulebohn



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