

**EQ:** Describe atomic structure.

- Atoms are made of extremely tiny particles called protons, neutrons, and electrons.
- Protons and neutrons are in the center of the atom, making up the nucleus.
- Electrons surround the nucleus.
- Protons have a positive charge.
- Electrons have a negative charge.
- The charge on the proton and electron are exactly the same size but opposite.
- Neutrons have no charge.
- Since opposite charges attract protons and electrons attract each other.

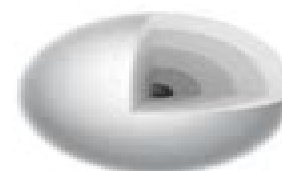
- ☐ Label the nucleus (protons, neutrons) and electrons in the drawing of a carbon atom.
- ☐ Draw a line between the subatomic particle and its charge.

proton  
electron  
neutron

no charge  
positive charge  
negative charge



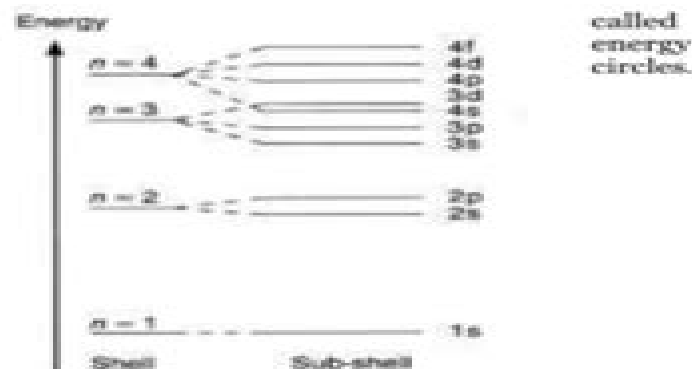
- The electrons surrounding an atom are located in regions around the nucleus called "energy levels".
- An energy level represents the 3-dimensional space surrounding the nucleus where electrons are most likely to be.
- The first energy level is closest to the nucleus. The second energy level is a little farther away than the first. The third is a little farther away than the second, and so on.
- Each energy level can accommodate or "hold" a different number of electrons before additional electrons begin to go into the next level.
- When the first energy level has 2 electrons, the next electrons go into the second energy level until the second level has 8 electrons.
- When the second energy level has 8 electrons, the next electrons go into the third energy level until the third level has 8 electrons.
- When the third energy level has 8 electrons, the next 2 electrons go into the fourth energy level.
- The electrons in the energy level farthest from the nucleus are called valence electrons.
- Atoms in the same column (group) in the periodic table have the same number of valence electrons.



## Energy levels

Electrons surround the nucleus of an atom in regions *energy levels*. Even though atoms are spherical, the levels in an atom are more easily shown in concentric

Energy Level	Maximum # electrons
1	2
2	8
3	18
4	32
5	50
6	72



# Atomic Structure Study Guide

**Arshad Iqbal**



## Atomic Structure Study Guide:

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and Sulfur Questions Chapter 21 Organic and Nitrogen Compounds Questions Chapter 22 Periodicity Questions Chapter 23 Polymerization Questions Chapter 24 Rates of Reaction Questions Chapter 25 Reaction Kinetics Questions Chapter 26 Redox Reactions and Electrolysis Questions Chapter 27 States of Matter Questions Chapter 28 Transition Elements Questions The Alcohols and Esters Quiz Questions PDF e Book Chapter 1 interview questions and answers on Introduction to alcohols and alcohols reactions The Atomic Structure and Theory Quiz Questions PDF e Book Chapter 2 interview questions and answers on Atom facts elements and atoms number of nucleons protons electrons and neutrons The Benzene Chemical Compound Quiz Questions PDF e Book Chapter 3 interview questions and answers on Introduction to benzene arenes reaction phenol and properties and reactions of phenol The Carbonyl Compounds Quiz Questions PDF e Book Chapter 4 interview questions and answers on Introduction to carbonyl compounds aldehydes and ketone testing nucleophilic addition with HCN preparation of aldehydes and ketone reduction of aldehydes and ketone The Carboxylic Acids and Acyl Compounds Quiz Questions PDF e Book Chapter 5 interview questions and answers on Acidity of carboxylic acids acyl chlorides ethanoic acid and reactions to form tri iodomethane The Chemical Bonding Quiz Questions PDF e Book Chapter 6 interview questions and answers on Chemical bonding types chemical bonding electron pair bond angle bond energy bond energy bond length bonding and physical properties bonding energy repulsion theory covalent bonding covalent bonds double covalent bonds triple covalent bonds electron pair repulsion and bond angles electron pair repulsion theory enthalpy change of vaporization intermolecular forces ionic bonding ionic bonds and covalent bonds ionic bonds metallic bonding metallic bonding and delocalized electrons number of electrons sigma bonds and pi bonds sigma bonds pi bonds s orbital and p orbital Van der Waals forces and contact points The Chemistry of Life Quiz Questions PDF e Book Chapter 7 interview questions and answers on Introduction to chemistry enzyme specificity enzymes reintroducing amino acids and proteins The Electrode Potential Quiz Questions PDF e Book Chapter 8 interview questions and answers on Electrode potential cells and batteries E Plimsoll values electrolysis process measuring standard electrode potential quantitative electrolysis redox and oxidation The Electrons in Atoms Quiz Questions PDF e Book Chapter 9 interview questions and answers on Electronic configurations electronic structure evidence ionization energy periodic table simple electronic structure sub shells and atomic orbitals The Enthalpy Change Quiz Questions PDF e Book Chapter 10 interview questions and answers on Standard enthalpy changes bond energies enthalpies Hess law introduction to energy changes measuring enthalpy changes The Equilibrium Quiz Questions PDF e Book Chapter 11 interview questions and answers on Equilibrium constant expression equilibrium position acid base equilibria chemical industry equilibria ethanoic acid gas reactions equilibria and reversible reactions The Group IV Quiz Questions PDF e Book Chapter 12 interview questions and answers on Introduction to group IV metallic character of group IV elements ceramic silicon oxide covalent bonds properties variation in group IV relative stability of oxidation states and tetra chlorides The Groups II and VII Quiz Questions PDF e Book Chapter 13 interview questions and answers on Atomic

number of group II metals covalent bonds density of group II elements disproportionation fluorine group II elements and reactions group VII elements and reactions halogens and compounds ionic bonds melting points of group II elements metallic radii of group II elements periodic table elements physical properties of group II elements physical properties of group VII elements reaction of group II elements with oxygen reactions of group II elements reactions of group VII elements thermal decomposition of carbonates and nitrates thermal decomposition of group II carbonates thermal decomposition of group II nitrates uses of group II elements uses of group II metals uses of halogens and their compounds The Halogenoalkanes Quiz Questions PDF e Book Chapter 14 interview questions and answers on Halogenoalkanes uses of halogenoalkanes elimination reactions nucleophilic substitution in halogenoalkanes and nucleophilic substitution reactions The Hydrocarbons Quiz Questions PDF e Book Chapter 15 interview questions and answers on Introduction to alkanes sources of alkanes addition reactions of alkenes alkane reaction alkenes and formulas The Introduction to Organic Chemistry Quiz Questions PDF e Book Chapter 16 interview questions and answers on Organic chemistry functional groups organic reactions naming organic compounds stereoisomerism structural isomerism and types of organic reactions The Ionic Equilibria Quiz Questions PDF e Book Chapter 17 interview questions and answers on Introduction to ionic equilibria buffer solutions equilibrium and solubility indicators and acid base titrations pH calculations and weak acids The Lattice Energy Quiz Questions PDF e Book Chapter 18 interview questions and answers on Introduction to lattice energy ion polarization lattice energy value atomization and electron affinity Born Haber cycle and enthalpy changes in solution The Moles and Equations Quiz Questions PDF e Book Chapter 19 interview questions and answers on Amount of substance atoms molecules mass chemical formula and equations gas volumes mole calculations relative atomic mass solutions and concentrations The Nitrogen and Sulfur Quiz Questions PDF e Book Chapter 20 interview questions and answers on Nitrogen gas nitrogen and its compounds nitrogen and gas properties ammonia ammonium compounds environmental problems caused by nitrogen compounds and nitrate fertilizers sulfur and oxides sulfuric acid and properties and uses of sulfuric acid The Organic and Nitrogen Compounds Quiz Questions PDF e Book Chapter 21 interview questions and answers on Amides in chemistry amines amino acids peptides and proteins The Periodicity Quiz Questions PDF e Book Chapter 22 interview questions and answers on Acidic oxides basic oxides aluminum oxide balancing equation period 3 chlorides balancing equations reactions with chlorine balancing equations reactions with oxygen bonding nature of period 3 oxides chemical properties of chlorine chemical properties of oxygen chemical properties periodicity chemistry periodic table chemistry oxides chlorides of period 3 elements electrical conductivity in period 3 oxides electronegativity of period 3 oxides ionic bonds molecular structures of period 3 oxides oxidation number of oxides oxidation numbers oxides and hydroxides of period 3 elements oxides of period 3 elements period III chlorides periodic table electronegativity physical properties periodicity reaction of sodium and magnesium with water and relative melting point of period 3 oxides The Polymerization Quiz Questions PDF e Book Chapter 23 interview questions

and answers on Types of polymerization polyamides polyesters and polymer deductions The Rates of Reaction Quiz Questions PDF e Book Chapter 24 interview questions and answers on Catalysis collision theory effect of concentration reaction kinetics and temperature effect on reaction rate The Reaction Kinetics Quiz Questions PDF e Book Chapter 25 interview questions and answers on Reaction kinetics catalysts kinetics and reaction mechanism order of reaction rare constant  $k$  and rate of reaction The Redox Reactions and Electrolysis Quiz Questions PDF e Book Chapter 26 interview questions and answers on Redox reaction electrolysis technique oxidation numbers redox and electron transfer The States of Matter Quiz Questions PDF e Book Chapter 27 interview questions and answers on states of matter ceramics gaseous state liquid state materials conservations and solid state The Transition Elements Quiz Questions PDF e Book Chapter 28 interview questions and answers on transition element ligands and complex formation physical properties of transition elements redox and oxidation

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## **Table of Contents Atomic Structure Study Guide**

1. Understanding the eBook Atomic Structure Study Guide
  - The Rise of Digital Reading Atomic Structure Study Guide
  - Advantages of eBooks Over Traditional Books
2. Identifying Atomic Structure Study Guide
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Atomic Structure Study Guide
  - User-Friendly Interface
4. Exploring eBook Recommendations from Atomic Structure Study Guide
  - Personalized Recommendations
  - Atomic Structure Study Guide User Reviews and Ratings
  - Atomic Structure Study Guide and Bestseller Lists
5. Accessing Atomic Structure Study Guide Free and Paid eBooks

- Atomic Structure Study Guide Public Domain eBooks
  - Atomic Structure Study Guide eBook Subscription Services
  - Atomic Structure Study Guide Budget-Friendly Options
6. Navigating Atomic Structure Study Guide eBook Formats
    - ePub, PDF, MOBI, and More
    - Atomic Structure Study Guide Compatibility with Devices
    - Atomic Structure Study Guide Enhanced eBook Features
  7. Enhancing Your Reading Experience
    - Adjustable Fonts and Text Sizes of Atomic Structure Study Guide
    - Highlighting and Note-Taking Atomic Structure Study Guide
    - Interactive Elements Atomic Structure Study Guide
  8. Staying Engaged with Atomic Structure Study Guide
    - Joining Online Reading Communities
    - Participating in Virtual Book Clubs
    - Following Authors and Publishers Atomic Structure Study Guide
  9. Balancing eBooks and Physical Books Atomic Structure Study Guide
    - Benefits of a Digital Library
    - Creating a Diverse Reading Collection Atomic Structure Study Guide
  10. Overcoming Reading Challenges
    - Dealing with Digital Eye Strain
    - Minimizing Distractions
    - Managing Screen Time
  11. Cultivating a Reading Routine Atomic Structure Study Guide
    - Setting Reading Goals Atomic Structure Study Guide
    - Carving Out Dedicated Reading Time
  12. Sourcing Reliable Information of Atomic Structure Study Guide
    - Fact-Checking eBook Content of Atomic Structure Study Guide
    - Distinguishing Credible Sources
  13. Promoting Lifelong Learning
    - Utilizing eBooks for Skill Development

- Exploring Educational eBooks

#### 14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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