#### Chapter 17: Acid-Base Equilibria

#### 17.2: The Buffer Solutions

Buffered Solution: - a solution that resists a change in pH when a small amount of H<sub>2</sub>O' or OH' is added. consists of a pair of weak acid/conjugate base common ion or a pair of weak

base/conjugate acid common ion.

- Acidic Buffered Solution:  $HA_{(aq)} + H_2O_{(0)} = H_2O^*_{(aq)} + A^*_{(aq)}$
- a. Small Amounts of H<sub>2</sub>O is Added:  $A^{-}_{(aa)} + H_1O^{+}_{(aa)} \rightarrow H_2O_{(0)} + H_{A(aa)}$ (H' - Strong Acid reacts completely with A')
  - (More HA Weak Acid: pH will only be lowered SLIGHTLY!)
- b. Small Amounts of OH is Added: HA tort + OH tort -> H2O to + A tort (OH" - Strong Base reacts completely with HA)
- (More A Weak Base: pH will only be raised SLIGHTLY!) B (ag) + H2O (b) = OH (ag) + HB (ag) Basic Buffered Solution:
- a. Small Amounts of HyO' is Added:  $\mathbf{B}_{(aq)} + \mathbf{H}_3\mathbf{O}^*_{(aq)} \rightarrow \mathbf{H}_2\mathbf{O}_{(0)} + \mathbf{H}\mathbf{B}^*_{(aq)}$
- (H' Strong Acid reacts completely with B)
- (More HB' Weak Acid: pH will only be lowered SLIGHTLY!)
- b. Small Amounts of OH<sup>-</sup> is Added: B (ag) + H2O (b) = OH" (ag) + HB" (ag) (OH - Strong Base increases [OH ]; eq shifts left) (More B – Weak Base: pH will only be raised SLIGHTLY!)
- Example: HCO<sub>1</sub> and CO<sub>2</sub> is a buffer because HCO<sub>1</sub> acts as an acid and CO<sub>2</sub> acts as a base whenever a small amount of strong base or acid is added to it. Thus, pH will remain somewhat constant.
  - HCO<sub>2</sub> (ast) + OH (ast) → H<sub>2</sub>O<sub>10</sub> + CO<sub>2</sub> (ast) When small amounts of strong base is added:  $CO_3^{2-}$  (and  $+ H_3O^+$  (and  $\rightarrow H_2O_4$ )  $+ HCO_3$  (and When small amounts of strong acid is added:
- Example 1: Specify which of these systems can be classified as a buffer system.
- a. HNO<sub>1</sub> / KNO<sub>1</sub> b. CH<sub>3</sub>COOH / NaCH<sub>3</sub>COO HNOs is a strong CH2COOH is a weak acid.
- acid. Hence, NO1 is Therefore, CH<sub>2</sub>COO is a not a strong enough strong enough weak base to base to resist the resist the change in pH. change in pH. HNO<sub>3</sub> NaCH3COO provides this / KNO<sub>1</sub> is NOT a buffer system.
  - base. CH<sub>3</sub>COOH / CH<sub>3</sub>COO is a buffer system.
- c. H<sub>2</sub>SO<sub>4</sub> / NaHSO<sub>4</sub> H-SO, is a strong acid. Hence, HSO<sub>4</sub> is not a strong enough base to resist the change in pH. H2SO4 / NaHSO4 is

NOT a buffer

NH<sub>3</sub> is a weak base. Therefore, NHL is a strong enough weak acid to resist the change in pH. NH<sub>2</sub>CI provides this acid. NHA/NHACL is a buffer system.

d. NH<sub>3</sub> / NH<sub>4</sub>CI

- Example 2: Calculate the pH of a 1.00 L buffered solution consisting of 0.500 M of methylamine  $(K_b = 4.38 \times 10^{-4})$  with 0.450 M of CH<sub>3</sub>NH<sub>3</sub>Cl when:
  - a. 0.0200 mol of KOH is added to it.
  - b. 3.00 mL of 0.750 M of HNO<sub>3 (m)</sub> is added to it.
  - c. Contrast the pH's of the above two additions against the same additions to 1.00 L of water.

# **Acid Baase Equilibria Study Guide Answers**

**Roman Wölfel** 

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