

7 Acid, bases and salts

7.1

Acids and bases

LEARNING OUTCOMES

At the end of this topic you should be able to:

- define acid, acid anhydride, alkali, base and salt
- explain acids and bases in terms of replaceable hydrogen and proton transfer.



Figure 7.1.1 All these household products contain acids.

EXAM TIP

The sign $\longrightarrow$ is used when the reaction goes to completion. For strong acids and bases, the reaction always goes to completion. The sign $\rightleftharpoons$ is used when the reaction does not go to completion and both reactants and products are present (see 7.2). This applies to weak acids and bases.

What are acids and bases?

Acids usually have a sour taste, change blue litmus to red, and are corrosive. The commonest inorganic acids are hydrochloric acid, HCl(aq) , sulfuric acid, $\text{H}_2\text{SO}_4\text{(aq)}$ and nitric acid, $\text{HNO}_3\text{(aq)}$. A simple definition of an acid is that it is a substance that neutralises a **base** to form a salt and water. For example:



Acids have one or more **replaceable hydrogen** atoms. The hydrogen is replaced by a metal or ammonium ion, NH_4^+ . In the reaction above, the hydrogen in the acid is replaced by the calcium ion.

- Monobasic acids have one replaceable hydrogen atom, e.g. HCl .
- Dibasic acids have two replaceable hydrogen atoms, e.g. sulfuric acid:



- Tribasic acids, e.g. H_3PO_4 , have three replaceable hydrogen atoms.

A **base** is a substance that neutralises an acid to form a salt and water. Bases are often oxides and hydroxides of metals.

A base that is soluble in water is called an **alkali**. The hydroxides of the Group I metals, e.g. NaOH , and aqueous ammonia, $\text{NH}_3\text{(aq)}$, are alkalis.

Salts

A **salt** is a compound formed when the hydrogen in an acid is replaced by a metal or ammonium ion.

- Chloride salts are formed from hydrochloric acid. For example, calcium chloride:



- Sulfates are formed from sulfuric acid. For example, sodium sulfate:



- Nitrates are formed from nitric acid. For example, potassium nitrate:



Proton transfer in acids and bases

When an acid dissolves in water, it ionises (splits up into ions). Hydrogen ions are formed. The equation below shows hydrogen chloride gas ionising in water (aq).



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Arshad Iqbal



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1. Understanding the eBook Acids Bases And Salts Study Guide Awnsers
 - The Rise of Digital Reading Acids Bases And Salts Study Guide Awnsers
 - Advantages of eBooks Over Traditional Books
2. Identifying Acids Bases And Salts Study Guide Awnsers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Acids Bases And Salts Study Guide Awnsers
 - User-Friendly Interface
4. Exploring eBook Recommendations from Acids Bases And Salts Study Guide Awnsers
 - Personalized Recommendations
 - Acids Bases And Salts Study Guide Awnsers User Reviews and Ratings
 - Acids Bases And Salts Study Guide Awnsers and Bestseller Lists

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 - Acids Bases And Salts Study Guide Awnsers Budget-Friendly Options
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 - ePub, PDF, MOBI, and More
 - Acids Bases And Salts Study Guide Awnsers Compatibility with Devices
 - Acids Bases And Salts Study Guide Awnsers Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Acids Bases And Salts Study Guide Awnsers
 - Highlighting and Note-Taking Acids Bases And Salts Study Guide Awnsers
 - Interactive Elements Acids Bases And Salts Study Guide Awnsers
8. Staying Engaged with Acids Bases And Salts Study Guide Awnsers
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Acids Bases And Salts Study Guide Awnsers
9. Balancing eBooks and Physical Books Acids Bases And Salts Study Guide Awnsers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Acids Bases And Salts Study Guide Awnsers
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Acids Bases And Salts Study Guide Awnsers
 - Setting Reading Goals Acids Bases And Salts Study Guide Awnsers
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Acids Bases And Salts Study Guide Awnsers
 - Fact-Checking eBook Content of Acids Bases And Salts Study Guide Awnsers
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
- Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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