

Solutions, acids, and bases

1. VOCABULARY REVIEW OF SOLUTIONS

Solute - what is **dissolved** in the solvent to form a solution.

Solvent - the substance that **dissolves** the solute, ordinarily a liquid when the solute and the solvent come together to form a **homogeneous mixture**.

Miscible - two mix in all proportions (none to be mixed) - liquids will pour each over

Immiscible - don't mix in any proportions (none to be mixed)

Examples of miscible/immiscible - ethanol/water/methanol, etc. oil/water/acetone

Solute - aspect of being **dissolved** in liquid

Insoluble - a substance (tends to be a solid) that **will not dissolve** in a solvent - an example is sand and water

Suspension - looks like a **solution** when **stirred** but **settles** a bit of water (suspensions) also it is a mixture - gaseous mixture where the small solid particles of a substance are spread throughout a liquid without settling.

Unsaturated - solutions with solubility chart when the temperature would be enough to make a substance dissolve but at a greater amount of solute solubility **less solute dissolves than can** (ex. salt @ 20°C temp)

Saturated - solutions with solubility chart when the temperature-solubility relationship is at the line in other words a solution with solute has **reached** at the maximum possible amount.

Supersaturated - contains **more than the max. amount** of solute that can be dissolved at the temperature on the solubility chart this would mean the grams of solute would go above the line.

2. SOLUBILITY VS. TEMPERATURE CHART



Above the line - **supersaturated**

At the line - **saturated**

Under the line - **unsaturated**

3. METHODS OF SEPARATION

Filtration - separate particles by **size**

Distillation - separate by **boiling point** - where two liquids to be separated.

Evaporation - where one to remove a **liquid from a solid** by **transforming** the liquid to a **gaseous state**.

Centrifugation - separating molecules by their **densities** by **spinning** them apart.

Decanting - process of separating insoluble or suspension solutions by **pouring** away the liquid from the solid.

Chromatography - separation by the **rates** at which components of different solutions **move through** a medium (ex. gel electrophoresis)

Acid And Bases Study Guide

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Table of Contents Acid And Bases Study Guide

1. Understanding the eBook Acid And Bases Study Guide
 - The Rise of Digital Reading Acid And Bases Study Guide
 - Advantages of eBooks Over Traditional Books
2. Identifying Acid And Bases Study Guide
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Acid And Bases Study Guide
 - User-Friendly Interface
4. Exploring eBook Recommendations from Acid And Bases Study Guide
 - Personalized Recommendations
 - Acid And Bases Study Guide User Reviews and Ratings
 - Acid And Bases Study Guide and Bestseller Lists
5. Accessing Acid And Bases Study Guide Free and Paid eBooks
 - Acid And Bases Study Guide Public Domain eBooks
 - Acid And Bases Study Guide eBook Subscription Services
 - Acid And Bases Study Guide Budget-Friendly Options

6. Navigating Acid And Bases Study Guide eBook Formats
 - ePub, PDF, MOBI, and More
 - Acid And Bases Study Guide Compatibility with Devices
 - Acid And Bases Study Guide Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Acid And Bases Study Guide
 - Highlighting and Note-Taking Acid And Bases Study Guide
 - Interactive Elements Acid And Bases Study Guide
8. Staying Engaged with Acid And Bases Study Guide
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Acid And Bases Study Guide
9. Balancing eBooks and Physical Books Acid And Bases Study Guide
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Acid And Bases Study Guide
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Acid And Bases Study Guide
 - Setting Reading Goals Acid And Bases Study Guide
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Acid And Bases Study Guide
 - Fact-Checking eBook Content of Acid And Bases Study Guide
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
 - Integration of Multimedia Elements

- Interactive and Gamified eBooks

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