

Acids, salts and bases

Acid reactions:

Acid + metal → salt + hydrogen

Sulphuric acid + iron → iron sulphate + hydrogen

Base/alkali + acid → salt + water

Sodium hydroxide + hydrochloric acid → Sodium chloride + water

Acid + metal carbonate → salt + carbon dioxide + water

Nitric acid + magnesium oxide → magnesium nitrate + carbon dioxide + water

Alkali + ammonium salt → salt + ammonia + water

Sodium hydroxide + ammonium chloride → sodium chloride + ammonia + water

Acid	Formulae		Formulae
Hydrochloric acid	HCl	Chloride	
Nitric acid	HNO ₃	Nitrate	
Sulphuric acid	H ₂ SO ₄	Sulphate	

	Name	Ions present
Acids	Hydrochloric acid	H ⁺ _(aq) and Cl ⁻ _(aq)
	Nitric acid	H ⁺ _(aq) and NO ₃ ⁻ _(aq)
	Sulphuric acid	H ⁺ _(aq) , HSO ₄ ⁻ _(aq) and SO ₄ ²⁻ _(aq)
Alkalis	Sodium hydroxide	Na ⁺ _(aq) and OH ⁻ _(aq)
	Potassium hydroxide	K ⁺ _(aq) and OH ⁻ _(aq)
	Calcium hydroxide	Ca ²⁺ _(aq) and OH ⁻ _(aq)
	Ammonia solution	NH ₄ ⁺ _(aq) and OH ⁻ _(aq)

Acids Bases And Salts Core Teaching Resources

Rachel Sandford



Acids Bases And Salts Core Teaching Resources:

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