

## 14.3

## HYDROGEN IONS AND ACIDITY

## Section Review

## Objectives

1. Calculate concentrations based on the pH of a given acid or hydrogen ion concentration.
2. Calculate hydrogen ion concentrations from  $\text{pH}$  and  $\text{pOH}$ , and vice versa from  $\text{pH}$  and  $\text{pOH}$ .
3. Calculate the  $\text{pH}$  of a weak acid.

## Vocabulary

- |                   |                               |                  |
|-------------------|-------------------------------|------------------|
| 1. autoionization | 2. hydrogen ion concentration | 3. acid solution |
| 4. autoionization | 5. $\text{pH}$                | 6. $\text{pOH}$  |

## Key Equation

- |                                     |                                  |
|-------------------------------------|----------------------------------|
| 1. $\text{pH} = -\log [\text{H}^+]$ | 2. $\text{pH} = 14 - \text{pOH}$ |
| 3. $\text{pH} = -\log [\text{H}^+]$ | 4. $\text{pH} = 14 - \text{pOH}$ |

## Part A Completion

Read the statement and fill in the blank with the word or phrase that is indicated in parentheses. Write the word or phrase in the space provided.

1. When a substance ionizes, it releases hydrogen ions. (acidic)
2. The concentration of  $\text{H}^+$  in the autoionization of water is  $1.0 \times 10^{-7}$  moles per liter at 25°C. (neutral)
3. The  $\text{pH}$  of a solution is the negative log of the concentration of  $\text{H}^+$ . (acidic)
4. The  $\text{pH}$  of a solution is the negative log of the concentration of  $\text{H}^+$ . (acidic)
5. The  $\text{pH}$  of a solution is the negative log of the concentration of  $\text{H}^+$ . (acidic)
6. The  $\text{pH}$  of a solution is the negative log of the concentration of  $\text{H}^+$ . (acidic)
7. The  $\text{pH}$  of a solution is the negative log of the concentration of  $\text{H}^+$ . (acidic)
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17. The  $\text{pH}$  of a solution is the negative log of the concentration of  $\text{H}^+$ . (acidic)
18. The  $\text{pH}$  of a solution is the negative log of the concentration of  $\text{H}^+$ . (acidic)
19. The  $\text{pH}$  of a solution is the negative log of the concentration of  $\text{H}^+$ . (acidic)
20. The  $\text{pH}$  of a solution is the negative log of the concentration of  $\text{H}^+$ . (acidic)

1.  $\text{H}^+$  concentration2.  $1.0 \times 10^{-7}$  M3.  $1.0 \times 10^{-7}$  M4.  $1.0 \times 10^{-7}$  M5.  $1.0 \times 10^{-7}$  M6.  $1.0 \times 10^{-7}$  M7.  $1.0 \times 10^{-7}$  M8.  $1.0 \times 10^{-7}$  M9.  $1.0 \times 10^{-7}$  M10.  $1.0 \times 10^{-7}$  M11.  $1.0 \times 10^{-7}$  M12.  $1.0 \times 10^{-7}$  M13.  $1.0 \times 10^{-7}$  M14.  $1.0 \times 10^{-7}$  M15.  $1.0 \times 10^{-7}$  M16.  $1.0 \times 10^{-7}$  M17.  $1.0 \times 10^{-7}$  M18.  $1.0 \times 10^{-7}$  M19.  $1.0 \times 10^{-7}$  M20.  $1.0 \times 10^{-7}$  M

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**David Kirk**



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## **Table of Contents 19 2 Hydrogen Ions And Acidity Work Sheet Answers**

1. Understanding the eBook 19 2 Hydrogen Ions And Acidity Work Sheet Answers
  - The Rise of Digital Reading 19 2 Hydrogen Ions And Acidity Work Sheet Answers
  - Advantages of eBooks Over Traditional Books
2. Identifying 19 2 Hydrogen Ions And Acidity Work Sheet Answers
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an 19 2 Hydrogen Ions And Acidity Work Sheet Answers
  - User-Friendly Interface
4. Exploring eBook Recommendations from 19 2 Hydrogen Ions And Acidity Work Sheet Answers
  - Personalized Recommendations
  - 19 2 Hydrogen Ions And Acidity Work Sheet Answers User Reviews and Ratings
  - 19 2 Hydrogen Ions And Acidity Work Sheet Answers and Bestseller Lists
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  - 19 2 Hydrogen Ions And Acidity Work Sheet Answers Budget-Friendly Options
6. Navigating 19 2 Hydrogen Ions And Acidity Work Sheet Answers eBook Formats

- ePub, PDF, MOBI, and More
  - 19 2 Hydrogen Ions And Acidity Work Sheet Answers Compatibility with Devices
  - 19 2 Hydrogen Ions And Acidity Work Sheet Answers Enhanced eBook Features
7. Enhancing Your Reading Experience
    - Adjustable Fonts and Text Sizes of 19 2 Hydrogen Ions And Acidity Work Sheet Answers
    - Highlighting and Note-Taking 19 2 Hydrogen Ions And Acidity Work Sheet Answers
    - Interactive Elements 19 2 Hydrogen Ions And Acidity Work Sheet Answers
  8. Staying Engaged with 19 2 Hydrogen Ions And Acidity Work Sheet Answers
    - Joining Online Reading Communities
    - Participating in Virtual Book Clubs
    - Following Authors and Publishers 19 2 Hydrogen Ions And Acidity Work Sheet Answers
  9. Balancing eBooks and Physical Books 19 2 Hydrogen Ions And Acidity Work Sheet Answers
    - Benefits of a Digital Library
    - Creating a Diverse Reading Collection 19 2 Hydrogen Ions And Acidity Work Sheet Answers
  10. Overcoming Reading Challenges
    - Dealing with Digital Eye Strain
    - Minimizing Distractions
    - Managing Screen Time
  11. Cultivating a Reading Routine 19 2 Hydrogen Ions And Acidity Work Sheet Answers
    - Setting Reading Goals 19 2 Hydrogen Ions And Acidity Work Sheet Answers
    - Carving Out Dedicated Reading Time
  12. Sourcing Reliable Information of 19 2 Hydrogen Ions And Acidity Work Sheet Answers
    - Fact-Checking eBook Content of 19 2 Hydrogen Ions And Acidity Work Sheet Answers
    - Distinguishing Credible Sources
  13. Promoting Lifelong Learning
    - Utilizing eBooks for Skill Development
    - Exploring Educational eBooks
  14. Embracing eBook Trends
    - Integration of Multimedia Elements
    - Interactive and Gamified eBooks

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