

- 6) In the first trial, the rate constant, k , obtained at 25°C is indicated in a spreadsheet along with a calculated value of k . What is the rate constant?

$$\begin{aligned}
 & \text{Reaction: } \text{CO}_2(g) + \text{H}_2\text{O}(l) \rightleftharpoons \text{H}^+(aq) + \text{HCO}_3^-(aq) \\
 & \text{Rate: } -\frac{d[\text{CO}_2]}{dt} = \frac{d[\text{H}^+]}{dt} = \frac{d[\text{HCO}_3^-]}{dt} \\
 & \text{Rate constant: } k = \frac{\text{Rate}}{[\text{CO}_2]} = \frac{1.5 \times 10^{-4} \text{ M s}^{-1}}{0.01 \text{ M}} = 1.5 \times 10^{-2} \text{ s}^{-1}
 \end{aligned}$$

- 7) Predict whether each change will shift the equilibrium left, right, or have no effect for the following equilibrium reaction:



- adding more CO_2 → CO_2 will shift HCO_3^-
- adding more H_2O → H_2O will shift HCO_3^-
- adding a catalyst to the reaction mixture → CO_2 will shift HCO_3^-
- removing HCO_3^- → HCO_3^- will shift HCO_3^-
- adding more H^+ → H^+ will shift HCO_3^-

- 8) Calculate ΔG° and ΔG for the reaction (first), and ΔG for the reaction (second).

$$\begin{aligned}
 & \text{Reaction: } \text{CO}_2(g) + \text{H}_2\text{O}(l) \rightleftharpoons \text{H}^+(aq) + \text{HCO}_3^-(aq) \\
 & \Delta G^\circ = -RT \ln K = - (8.314 \text{ J K}^{-1} \text{ mol}^{-1}) (298 \text{ K}) \ln (1.5 \times 10^{-2}) = 1.5 \times 10^4 \text{ J mol}^{-1} \\
 & \Delta G = \Delta G^\circ + RT \ln Q = 1.5 \times 10^4 \text{ J mol}^{-1} + (8.314 \text{ J K}^{-1} \text{ mol}^{-1}) (298 \text{ K}) \ln (1.5 \times 10^{-2}) = 1.5 \times 10^4 \text{ J mol}^{-1}
 \end{aligned}$$

- 9) Calculate the standard Gibbs free energy of formation for each compound, and the standard Gibbs free energy of formation for the reaction.

$$\begin{aligned}
 & \text{Reaction: } \text{CO}_2(g) + \text{H}_2\text{O}(l) \rightleftharpoons \text{H}^+(aq) + \text{HCO}_3^-(aq) \\
 & \Delta G^\circ_f(\text{CO}_2) = -394.4 \text{ kJ mol}^{-1} \\
 & \Delta G^\circ_f(\text{H}_2\text{O}) = -237.1 \text{ kJ mol}^{-1} \\
 & \Delta G^\circ_f(\text{H}^+) = 0 \text{ kJ mol}^{-1} \\
 & \Delta G^\circ_f(\text{HCO}_3^-) = -689.8 \text{ kJ mol}^{-1}
 \end{aligned}$$

- 10) Calculate all the following results, using only the appropriate formula, constant, and the following data (given):

$$\begin{aligned}
 & \text{Reaction: } \text{CO}_2(g) + \text{H}_2\text{O}(l) \rightleftharpoons \text{H}^+(aq) + \text{HCO}_3^-(aq) \\
 & \Delta G^\circ_f(\text{CO}_2) = -394.4 \text{ kJ mol}^{-1} \\
 & \Delta G^\circ_f(\text{H}_2\text{O}) = -237.1 \text{ kJ mol}^{-1} \\
 & \Delta G^\circ_f(\text{H}^+) = 0 \text{ kJ mol}^{-1} \\
 & \Delta G^\circ_f(\text{HCO}_3^-) = -689.8 \text{ kJ mol}^{-1}
 \end{aligned}$$

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Linda Swisher, Kevin T. Patton



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chemistry also follows from the simple to the more complex and the strength of this text is that it enables students with varied backgrounds to proceed together to significant levels of achievement

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Basic Concepts of Chemistry, 9e Study Guide and Solutions Manual Leo J.

Malone, Theodore O. Dolter, 2012-01-03 The 9th edition of Malone's Basic Concepts of Chemistry provides many new and advanced features that continue to address general chemistry topics with an emphasis on outcomes assessment New and advanced features include an objectives grid at the end of each chapter which ties the objectives to examples within the sections assessment exercises at the end each section and relevant chapter problems at the end of each chapter A new Math Check allows quick access to the needed basic skill The first chapter now includes brief introductions to several fundamental chemical concepts and Chapter Synthesis Problems have been added to the end of each chapter to bring key concepts into one encompassing problem Every concept in the text is clearly illustrated with one or more step by step examples Making it Real essays have been updated to present timely and engaging real world applications emphasizing the relevance of the material they are learning This edition continues the end of chapter Student Workshop activities to cater to the many different learning styles and to engage users in the practical aspect of the material discussed in the chapter

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