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## **NATIONAL CERTIFICATE ENGINEERING DRAWING N2**

(8090272)

**29 November 2019 (X-Paper)  
09:00–13:00**

*This question paper consists of 10 pages.*

# N2 Drawing Previous Question Papers

**Marc van Kreveld, Bettina Speckmann**



## **N2 Drawing Previous Question Papers:**

*Graph Drawing* Guiseppe Liotta, 2004-03-04 The 11th International Symposium on Graph Drawing GD 2003 was held on September 21-24 2003 at the Università degli Studi di Perugia Perugia Italy GD 2003 attracted 93 participants from academic and industrial institutions in 17 countries In response to the call for papers the program committee received 88 regular submissions describing original research and/or system demonstrations Each submission was reviewed by at least 4 program committee members and comments were returned to the authors Following extensive e-mail discussions the program committee accepted 34 long papers 12 pages each in the proceedings and 11 short papers 6 pages each in the proceedings Also 6 posters 2 pages each in the proceedings were displayed in the conference poster gallery In addition to the 88 submissions the program committee also received a submission of special type one that was not competing with the others for a time slot in the conference program and that collects selected open problems in graph drawing The aim of this paper which was refereed with particular care and UNCHANGED two rounds of revisions is to stimulate future research in the graph drawing community The paper presents 42 challenging open problems in different areas of graph drawing and contains more than 120 references Although the length of the paper makes it closer to a journal version than to a conference extended abstract we decided to include it in the conference proceedings so that it could easily reach in a short time the vast majority of the graph drawing community

**Graph Drawing** Marc van Kreveld, Bettina Speckmann, 2011-12-16 This volume constitutes the refereed proceedings of the 19th International Symposium on Graph Drawing GD 2010 held in Eindhoven The Netherlands during September 2011 The 34 revised full papers presented together with 3 revised short and 6 poster papers were carefully reviewed and selected from 88 submissions Furthermore the proceedings contain the abstracts of two invited talks and to commemorate Kozo Sugiyama and his pioneering research in graph drawing the proceedings include an obituary A unique and fun part of the symposium is the Graph Drawing Contest which is part of the Graph Drawing Challenge This year was the 18th edition A report on the contest is included at the end of the proceedings

**Graph Drawing** Christian Duncan, Antonios Symvonis, 2014-11-26 This book constitutes the proceedings of the 22nd International Symposium on Graph Drawing GD 2014 held in Würzburg Germany in September 2014 The 41 full papers presented in this volume were carefully reviewed and selected from 72 submissions The back matter of the book also contains 2 page poster papers presented at the conference The contributions are organized in topical sections named planar subgraphs simultaneous embeddings applications contact representations k planar graphs crossing minimization level drawings theory fixed edge directions drawing under constraints clustered planarity and greedy graphs (Free Sample)

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Graph Drawing Petra Mutzel,Michael Jünger,Sebastian Leipert,2002-02-27 This book constitutes the thoroughly refereed post proceedings of the 9th International Symposium on Graph Drawing GD 2001 held in Vienna Austria in September 2001 The 32 revised full papers presented were carefully reviewed and selected from 66 paper submissions Also included are a corrected version of a paper from the predecessor volume short reports on the software systems exhibition two papers of the special session on graph exchange formats and a report on the annual graph drawing contests The papers are organized in topical sections on hierarchical drawing planarity crossing theory compaction planar graphs symmetries interactive drawing representations aesthetics 2D and 3D embeddings data visualization floor planning and planar drawing JEE Advanced 14 Year-wise Solved Papers 1 & 2 (2006 - 2019) Disha Experts,2019-07-02 JEE Advanced 14 Year wise Solved Papers 1 The book

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web 1 consider the reaction  $\text{I}_2\text{O}_5 + 5\text{CO} \rightarrow \text{I}_2 + 5\text{CO}_2$  80.0 grams of iodine(V) oxide reacts with 28.0 grams of carbon monoxide. CO determine the mass of iodine ( $\text{I}_2$ ) which could be produced. b) if in the above situation only 0.160 moles of iodine ( $\text{I}_2$ ) was produced, what mass of iodine was produced?

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web limiting reactant theoretical and percentage yields key  $4\text{KO}_2 + 2\text{H}_2\text{O} \rightarrow 4\text{KOH} + 3\text{O}_2$  if a reaction vessel contains 0.1 mol  $\text{KO}_2$  and 0.1 mol  $\text{H}_2\text{O}$ , what is the limiting reactant?

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web chem 12 3 limiting reagent and percent yield 4.06 reviews whenever quantities of two or more reactants are given in a stoichiometric problem, you must identify the limiting reagent. click the card to flip limiting reagent.

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web limiting reactant: reactant present in an amount lower than required by the reaction stoichiometry, thus limiting the amount of product generated. percent yield: measure of the efficiency of a reaction, expressed as a percentage of the theoretical yield.

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web limiting reagent stoichiometry google classroom you might need calculator periodic table given the following reaction:  
 $\text{Cu} + 2\text{AgNO}_3 \rightarrow \text{Cu}(\text{NO}_3)_2 + 2\text{Ag}$

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 $3\text{CO} + 7\text{H}_2 \rightarrow \text{C}_3\text{H}_8 + 3\text{H}_2\text{O}$

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web start with 8.3 mole  $\text{H}_2\text{S}$  and multiply it with 2 mole  $\text{H}_2\text{O}$  over 2 mole  $\text{H}_2\text{S}$  from the equation  $8.3 \text{ mole } \text{H}_2\text{S} \times \frac{2 \text{ mole } \text{H}_2\text{O}}{2 \text{ mole } \text{H}_2\text{S}}$



mole  $\text{H}_2\text{S}$  x 18g  $\text{H}_2\text{O}$  i mole  $\text{H}_2\text{O}$  149 4g  $\text{H}_2\text{O}$  is your theoretical yield so percent yield  $\frac{137}{149} \times 100 = 91.77\%$  problems 6 8 are similar to problem 5

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web 1 csun edu hcchm001 introchemhandouts html limiting reagents theoretical actual and percent yields example of a limiting reagent problem how many grams of  $\text{NH}_3$  can be produced theoretically from the reaction of 5.0 g of N what is the limiting reagent if 8.52 g are actually formed what is the

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web limiting reagents problem one 1 for the balanced equation  $\text{C}_3\text{H}_6 + 3\text{O}_2 \rightarrow 3\text{CO} + 3\text{H}_2\text{O}$  if the reaction of 21.3 grams of  $\text{O}_2$  produces a 47.3% yield how many grams of CO would be produced step one figure out how many moles of the limiting reagent you have  $\frac{21.3}{32} = 0.666$  moles  $\text{O}_2$  0.666 moles  $\text{CO}$  32 g/mol

experiment 4 stoichiometry limiting reagents yield making - Feb 26 2022

web percent yield is a measure of the efficiency of the experimental design yield efficiency  $\frac{\text{mass of product obtained}}{\text{calculated mass of the product expected}} \times 100$  in this reaction a yield 80 is anticipated

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web a write a balanced equation for the reaction  $\text{Al} + \text{O}_2 \rightarrow \text{Al}_2\text{O}_3$  4 Al + 3  $\text{O}_2 \rightarrow 2 \text{Al}_2\text{O}_3$  b determine the theoretical yield of  $\text{Al}_2\text{O}_3$  160.0 g Al x  $\frac{1 \text{ mol Al}}{27 \text{ g Al}} = 5.93$  mol  $\text{Al}_2\text{O}_3$  x 101.964 g  $\text{Al}_2\text{O}_3$  302.3 g  $\text{Al}_2\text{O}_3$  26.982 g Al 4 mol Al 1 mol  $\text{Al}_2\text{O}_3$  c determine the percent yield  $\frac{260.0}{302.3} \times 100 = 86.01\%$  theoretical 302.3 g

**7 2 limiting reagent and reaction yields chem 1114** - Jul 02 2022

web 7 Na + 2  $\text{C}_2\text{O}_4$  is the limiting reactant percent yield 86.6% 8 only four molecules can be made 9 this amount cannot be weighted by ordinary balances and is worthless 10 nitrogen is the limiting reagent 11 yes methane is the limiting reagent 12 C is the limiting reagent 4 33 g of  $\text{H}_2$  are left over 13

6 5 limiting reagent and percent yield chemistry libretexts - Aug 15 2023

web learning objectives define and determine theoretical yields actual yields and percent yields identify a limiting reagent from a set of reactants calculate how much product will be produced from the limiting reagent calculate how much reactant s remains when the reaction is complete

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web the percent yield is calculated as follows  $\text{percent yield} = \frac{\text{actual yield}}{\text{theoretical yield}} \times 100$  based on this definition we would expect a percent yield

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web may 20 2018 learning objectives identify the limiting reactant limiting reagent in a given chemical reaction calculate

how much product will be produced from the limiting reactant calculate how much reactant s remains when the reaction is complete

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web limiting reagent and percent yield practice worksheet answer key limiting reactant yield practice worksheet 1 methanol  $\text{CH}_3\text{OH}$  can be produced by the following reaction  $2\text{H}_2 + \text{CO} \rightarrow \text{CH}_3\text{OH}$  calculate the theoretical yield of  $\text{CH}_3\text{OH}$  if 68.5 g of CO is reacted with 8.6 g of  $\text{H}_2$

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web jun 2 2020 the percent yield of a reaction is the ratio of the actual yield to the theoretical yield multiplied by 100 to give a percentage text percent yield text actual yield g over text theoretical yield g times 100 label 3 7 3

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