

## **Prototype Summary: Ipratropium**

**Indications:** Maintenance treatment of bronchospasm associated with chronic obstructive pulmonary disease; treatment of seasonal allergic rhinitis as a nasal spray.

**Actions:** Anticholinergic that blocks vagally mediated reflexes by antagonizing the action of acetylcholine.

### **Pharmacokinetics:**

<u>Route</u>	<u>Onset</u>	<u>Peak</u>	<u>Duration</u>
Inhalation	15 min	1–2 h	3–4 h

**$T_{1/2}$ :** Unknown, metabolized by neural pathways.

**Adverse effects:** Nervousness, dizziness, headache, nausea, gastrointestinal distress, cough, palpitations.

# Quizlet 54 Hoursfor Pharmacology

**L Cohen**



**Quizlet 54 Hoursfor Pharmacology:**  
Pharmacology Review ,1966

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and put in series a small resistance as required from the design for instance for e1 you will need a battery with v 6v and you will need to put in series a small resistance of  $0.5\ \Omega$  to simulate the small resistance in a battery

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web 1 practice problem balance the redox reaction in basic solution  $\text{NO}_2 + \text{Al} \rightarrow \text{NH}_3 + \text{AlO}_2^{-}$  2 practice problem balance the following reduction oxidation reaction in basic solution what is the coefficient in front of  $\text{OH}^{-}$   $\text{Br}_2 + \text{I}^{-} \rightarrow \text{Br}^{-} + \text{I}_2$  previous topic next topic

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web google classroom for the following redox equation  $\text{MnO}_4^{-} + \text{SO}_2 + 2\text{H}^{+} \rightarrow \text{Mn}^{2+} + \text{H}_2\text{SO}_4$  what will be the stoichiometric coefficient of  $\text{H}_2\text{SO}_4$  if the above equation is balanced in an acidic medium stuck use a hint report a problem do 4 problems

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web balance this redox reaction by the half reaction method by assuming an acidic solution  $\text{Cr}_2\text{O}_7^{2-} + \text{Fe} \rightarrow \text{Cr}^{3+} + \text{Fe}^{3+}$  balance the redox reaction in exercise 9 by the half reaction method by assuming a basic solution

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web solution 1 the two half reactions balanced as if in acidic solution  $2\text{CN}^{-} + \text{Au} \rightarrow \text{Au}(\text{CN})_2^{-} + 2\text{e}^{-}$   $2\text{H}_2\text{O}_2 \rightarrow 2\text{H}_2\text{O} + \text{O}_2$  make electrons equal convert to basic solution  $4\text{CN}^{-} + 2\text{Au} \rightarrow 2\text{Au}(\text{CN})_2^{-} + 2\text{e}^{-}$  multiplied by a factor of 2  $2\text{H}_2\text{O}_2 \rightarrow 2\text{H}_2\text{O} + \text{O}_2$   $2\text{OH}^{-}$  the final answer

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web aug 29 2023 the first step to balance any redox reaction is to separate the reaction into half reactions the substance being reduced will have electrons as reactants and the oxidized substance will have electrons as products usually all reactions are written as reduction reactions in half reaction tables

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web we can balance it for charge by adding two electrons to the right side of the equation so that the net charge on each side is 0 oxidation  $\text{Ni} \rightarrow \text{Ni}^{2+} + 2\text{e}^-$  now that the oxidation half reaction is balanced it tells us that two electrons are produced for every atom of nickel oxidized but where do those electrons go

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web balancing redox reactions in acidic solution problems 1 10 problem 1  $\text{Cr}_2\text{O}_7^{2-} + \text{Fe}^{2+} + \text{Cr}^{3+} + \text{Fe}^{3+}$  solution 1 balanced half reactions  $6\text{e}^- + 14\text{H}^+ + \text{Cr}_2\text{O}_7^{2-} \rightarrow 2\text{Cr}^{3+} + 7\text{H}_2\text{O}$   $\text{Fe}^{2+} \rightarrow \text{Fe}^{3+} + \text{e}^-$  2 equalize the electrons  $6\text{e}^- + 14\text{H}^+ + \text{Cr}_2\text{O}_7^{2-} \rightarrow 2\text{Cr}^{3+} + 7\text{H}_2\text{O}$   $6\text{Fe}^{2+} \rightarrow 6\text{Fe}^{3+} + 6\text{e}^-$  multiplied by a factor of 6

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