



## Lab -Stoichiometry of Magnesium Oxide

*Goggles must be worn at all times*

### Introduction:

In this experiment, you will examine the reaction between magnesium metal and oxygen gas. When heated, magnesium reacts readily with oxygen in the air, to produce magnesium oxide. The chemical reaction is written as follows:  $2\text{Mg} + \text{O}_2 \rightarrow 2\text{MgO}$

The magnesium will be heated strongly in an open crucible for several minutes. You may need to repeat this step several times in order to ensure that all the magnesium has reacted and turned into product.

You will use stoichiometry to determine how much product you should be able to make from the experiment and then in doing the lab, you will see if you are actually able to produce the desired amount of  $\text{MgO}$ .

**Purpose:** 1) You will measure the mass of the magnesium that reacts and predict the mass of magnesium oxide that should be produced using stoichiometry.  
2) After the experiment you will calculate your % yield and % error of actual magnesium oxide produced.

### Materials Provided:

-Crucible	-Magnesium ribbon	-Bunsen burner	-Clay triangle
-Crucible tongs	-Ring stand w/ ring	-Safety goggles	

### Procedure:

- 1) Obtain a piece of magnesium ribbon from your teacher.
- 2) Obtain a clean, dry crucible. Find the mass of the crucible and record it on the Data Table.
- 3) Roll the magnesium in a loose coil, and place it in the crucible. Find the mass of the crucible, and magnesium. Record this mass on the Data Table.
- 4) Set up the ring stand, ring, Bunsen burner, and triangle. Place the crucible on the triangle.
- 5) Begin heating the crucible gradually. Heat slowly by moving the flame around underneath the crucible. Remove the heat temporarily if a large amount of smoke comes out of the crucible.
- 6) Heat the crucible to redness, for four minutes
- 7) Turn off the Bunsen burner and allow the crucible to cool to a temperature low enough so that you can touch the crucible.
- 8) Find the mass of the crucible & contents. Record this mass on the Data Table.
- 9) Reheat the crucible and contents for four minutes, cool, and find the mass again. Record this mass on the Data Table. Compare the masses found in steps 8 and 9. If the two masses are not the same, repeat this step again until the masses are the same.

# Magnesium Oxide Lab Answers

**Siddappa N.Byrareddy**



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