

- 15) How many grams are in 4.5 moles of sodium fluoride, NaF?

$$4.5 \text{ MOL NaF} \left(\frac{23 \text{ g Na} + 19 \text{ g F}}{1 \text{ MOL NaF}} \right) = \boxed{189 \text{ g NaF}}$$

- 16) How many moles are in 98.3 grams of aluminum hydroxide, Al(OH)_3 ?

$$98.3 \text{ g Al(OH)}_3 \cdot \left(\frac{1 \text{ MOL Al(OH)}_3}{27 \text{ g Al} + 16(3) \text{ g O} + 1(3) \text{ g H}} \right) = \frac{98.3 \text{ g}}{78 \text{ g}} = \boxed{1.26 \text{ MOL Al(OH)}_3}$$

- 17) How many grams are in 0.02 moles of beryllium iodide, BeI_2 ?

$$0.02 \text{ MOL BeI}_2 \left(\frac{9 \text{ g Be} + 127(2) \text{ g I}}{1 \text{ mole BeI}_2} \right) = \boxed{5.26 \text{ g BeI}_2}$$

- 18) How many moles are in 68 grams of copper (II) hydroxide, Cu(OH)_2 ?

$$68 \text{ g Cu(OH)}_2 \left(\frac{1 \text{ mol Cu(OH)}_2}{64 \text{ g Cu} + 16(2) \text{ g O} + 1(2) \text{ g H}} \right) = \boxed{0.694 \text{ mol Cu(OH)}_2}$$

- 19) How many grams are in 3.3 moles of potassium sulfide, K_2S ?

$$3.3 \text{ mol K}_2\text{S} \left(\frac{39(2) \text{ g K} + 32 \text{ g S}}{1 \text{ MOL K}_2\text{S}} \right) = \boxed{113.3 \text{ g K}_2\text{S}}$$

- 20) How many moles are in 1.2×10^3 grams of ammonia, NH_3 ?

$$1.2 \times 10^3 \text{ g NH}_3 \left(\frac{1 \text{ mol NH}_3}{14 + 1(3) \text{ g NH}_3} \right) = \boxed{70.6 \text{ MOL NH}_3}$$

- 21) How many grams are in 2.3×10^{-4} moles of calcium phosphate, $\text{Ca}_3(\text{PO}_4)_2$?

$$2.3 \times 10^{-4} \text{ MOL Ca}_3(\text{PO}_4)_2 \left(\frac{40(3) + 16(6) + 31(2) \text{ g Ca}_3(\text{PO}_4)_2}{1 \text{ mol Ca}_3(\text{PO}_4)_2} \right) = \boxed{0.06394 \text{ g Ca}_3(\text{PO}_4)_2}$$

- 22) How many moles are in 3.4×10^{-7} grams of silicon dioxide, SiO_2 ?

$$3.4 \times 10^{-7} \text{ MOL SiO}_2 \left(\frac{1 \text{ MOL SiO}_2}{28 + 16(2)} \right) = \boxed{5.67 \times 10^{-7} \text{ g SiO}_2}$$

- 23) How many grams are in 1.11 moles of manganese sulfate, $\text{Mn}_3(\text{SO}_4)_7$?

$$1.11 \text{ MOL Mn}_3(\text{SO}_4)_7 = \left(\frac{55(3) \text{ g Mn} + 32(7) \text{ g S} + 16(28) \text{ g O}}{1 \text{ mol Mn}_3(\text{SO}_4)_7} \right) = \frac{837}{1}$$

$$1.11 \text{ MOL Mn}_3(\text{SO}_4)_7 \left(\frac{837 \text{ g Mn}_3(\text{SO}_4)_7}{1 \text{ mol Mn}_3(\text{SO}_4)_7} \right) = \boxed{929.07 \text{ MOL Mn}_3(\text{SO}_4)_7}$$

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Michael Seilmaier



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