How many grams are in 4.5 moles of sodium fluoride, NaF?
$$4.5 \, \text{MOL NoF} \left(\frac{239 \, \text{No.} + 119 \, \text{F}}{4 \, \text{NoL NoF}} \right) = 189 \, \text{g No.F}$$

16) How many moles are in 98.3 grams of aluminum hydroxide, Al(OH)₃?
$$98.39 \text{ Al}(OH)_3 * \left(\frac{1 \text{ MOL} \text{ Al}(OH)_5}{27 \text{ghL} + 16(8) \text{gO} + 1(8) \text{gH}}\right) = \frac{98.3 \text{ g/m}}{78 \text{g}} = 1.26 \text{ MO}$$

17) How many grams are in 0.02 moles of beryllium iodide, Bel₂?

O.02 Mol Bel₂
$$\left(\frac{9\% \text{Be} + 127(2) \text{g I}}{\text{I Mole Bel}_2}\right) = \left[5.26 \text{ g Be}/z\right]$$

MOL

18) How many moles are in 68 grams of copper (II) hydroxide,
$$Cu(OH)_2$$
?

$$68 g Cu(OH)_2 \left(\frac{1}{64gCu + 16(2)gO + 1(2)gH} \right) = 0.694 \text{ mol}$$

$$Cu(OH)_2 \left(\frac{1}{64gCu + 16(2)gO + 1(2)gH} \right) = 0.694 \text{ mol}$$

How many moles are in
$$1.2 \times 10^3$$
 grams of ammonia, NH₃?
 1.2×10^3 g NH₃ $\left(\frac{4 \text{ mol NH3}}{14 + 10^3 \text{ grams}}\right) = 70.6 \text{ Mol NH3}$

How many moles are in 3.4 x 10⁷ grams of silicon dioxide, SiO₂?
$$3.4 \times 10^{-7} \text{ Mol. SiO2} \left(\frac{1 \text{ Mol. SiO2}}{28 + 16(z)} \right) = 5.67 \times 10^{-7} \text{ SiO2}$$

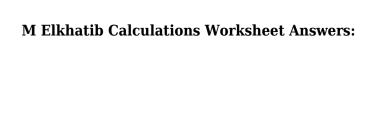
23) How many grams are in 1.11 moles of manganese sulfate,
$$Mn_3(SO_4)_7$$
?

||| $MOL Mn_3(SO_4)_7 = \frac{(55(3))7Mn + 32(7)95 + 16(28)40}{1 mol Mn_3(SO_4)^7} = \frac{837}{1}$

||| $MOL Mn_3(SO_4)_7 = \frac{(8379)7Mn_3(SO_4)_7}{1 mol Mn_3(SO_4)_7} = \frac{929.07}{100L}$

M Elkhatib Calculations Worksheet Answers

Michael Seilmaier



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