



5. (10 points)

The diagram above of pressure P versus volume V shows the expansion of 2.0 moles of a monatomic ideal gas from state A to state B. As shown in the diagram, $P_A = P_B = 600 \text{ N/m}^2$, $V_A = 3.0 \text{ m}^3$, and $V_B = 9.0 \text{ m}^3$.

(a)

- i. Calculate the work done by the gas as it expands.
- ii. Calculate the change in internal energy of the gas as it expands.
- iii. Calculate the heat added to or removed from the gas during this expansion.
- (b) The pressure is then reduced to 200 N/m² without changing the volume as the gas is taken from state B to state C. Label state C on the diagram and draw a line or curve to represent the process from state B to state C.
- (c) The gas is then compressed isothermally back to state A.
 - i. Draw a line or curve on the diagram to represent this process.
 - ii. Is heat added to or removed from the gas during this isothermal compression?

___ added to ____ removed from

Justify your answer.

14 Ap Physics B Response Questions Answers

Greg Jacobs

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