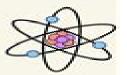


> ATOMIC STRUCTURE 2



Attorne structure -

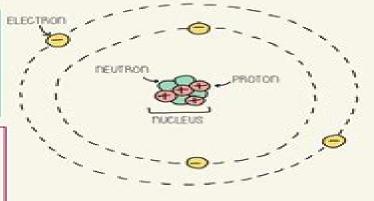
Every oform is made up of a central nucleus surrounded by electrons whitting around the nucleus in electron arbitals. The nucleus is they in relation to the size of the whole atom - comparable to a pea in the middle of a faotbal stadium. The nucleus is made up of profate and neutrons, which are the heavest particles. Bestrons are approximately 2000 times lighter than a proton or restron and weigh procladly nothing, therefore most of the mass comes from the tiny nucleus Profate are post-vely sharged whereas resultants, have no charge, making the nucleus positively charged overall Electrons are negatively charged, which enables the nucleus to keep haid of its electrons, due to the attraction between the apposite charges.

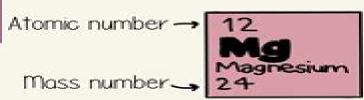
- Allomo number & mass number -

Each atom has an atomic number and a mass number, which are displayed on the Periodic Table as the small number and larger number next to each chemical symbol. The latomic number lets us the number of protons which is unique for each element. If the number of protons changes land therefore the atomic number shanges), which could occur during radioactive decay, for example, then the identity of the element also changes.

The mass number stells us the number of both protons and neutrons in the nucleus. Remember that most of an atom's mass comes from the nucleus and that the nucleus is made up of both protons and neutrons. Two atoms of the same element can have different mass numbers because they are isotopes.

Atomic particle	Charge	Mass
Proton	-1	
Deutron	0	
Electron		1/2000





- Isotopes -

befores are aforms which have the same number of profined but different numbers of newhold. Therefore they have the same atoms, number but a different mass number. For example, the most common safepes of carbon are carbon-12 and carbon-13, in which carbon-13 has an additional newtron in its nucleus to give it a higher mass number.

The mass numbers that we see in the periodic table are an average of all the masses of all satapes of a particular element. We call this the install we allow a formula and can be calculated using the following formula:

(Mass number of isotope 1 x percentage of isotope 1) • (mass number of isotope 2 x percentage of isotope 2)

Worked example:

Lets say we have a sample of carbon which consists of 76% of the solope carbon-12 and 24% carbon-13. Then the relative atomic mass (RATI) would be

RAM - (76 x 12) - (24 x 13) 100

BATS - 1229

Questions

- I flame the two subolamic particles that are found in the nucleus and state their charges
- 2. What does atomic number represent?
- 3. Why do atoms have an overall neutral charge?
- 4 Sadam Dalihas on atomic number of 11 and a mass number of 23 State. The number or profoss, neutrons and electrons in a sadam atom.
- 5 What is an isotope?"
- 6. A sample of bramine contains 45 / bramine-79 and 55 / bramine-81. Cabulate its relative atomic mass (RAMI)

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Kathryn Wolff,Susan M.
O'Connell,Valerie J. Montenegro

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 - The Rise of Digital Reading Atomic Structure Study Guide Answers Holt Science
 - Advantages of eBooks Over Traditional Books
- 2. Identifying Atomic Structure Study Guide Answers Holt Science
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
- 3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Atomic Structure Study Guide Answers Holt Science
 - User-Friendly Interface
- 4. Exploring eBook Recommendations from Atomic Structure Study Guide Answers Holt Science
 - Personalized Recommendations
 - Atomic Structure Study Guide Answers Holt Science User Reviews and Ratings
 - Atomic Structure Study Guide Answers Holt Science and Bestseller Lists
- 5. Accessing Atomic Structure Study Guide Answers Holt Science Free and Paid eBooks
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- Atomic Structure Study Guide Answers Holt Science Enhanced eBook Features
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 - Interactive Elements Atomic Structure Study Guide Answers Holt Science
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 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Atomic Structure Study Guide Answers Holt Science
- 9. Balancing eBooks and Physical Books Atomic Structure Study Guide Answers Holt Science
 - Benefits of a Digital Library
 - o Creating a Diverse Reading Collection Atomic Structure Study Guide Answers Holt Science
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Atomic Structure Study Guide Answers Holt Science
 - Setting Reading Goals Atomic Structure Study Guide Answers Holt Science
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Atomic Structure Study Guide Answers Holt Science
 - Fact-Checking eBook Content of Atomic Structure Study Guide Answers Holt Science
 - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - $\circ\,$ Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
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