

Atoms & The Periodic Table

Unit 3 Test Study Guide

Lesson 1: The Atom

1. Who was the ancient Greek philosopher that came up with the idea that all things are made of small "uncuttable" particles?
2. What is an atom?
3. What was John Dalton's contribution to the atomic theory?
4. What are the parts of an atom?
 - a. Where can each of these particles be found within an atom?
 - b. How do the electrical charges of protons, neutrons, and electrons compare to each other?
 - c. How do the masses of protons, neutrons, and electrons compare to each other?
5. What is an atomic number?
 - a. What does this tell about the structure of an atom?
6. What is the atomic mass (mass number)?
 - a. What does this tell about the structure of an atom?
 - i. How can the atomic mass help you find the number of neutrons in an atom?

Lesson 2: The Periodic Table

1. What are elements?
2. What is the Periodic Table of Elements?
 - a. How are elements arranged on the Periodic Table?
3. Information found on each square of the Periodic Table:
 - a. Atomic Number (*See Lesson 1*)
 - b. What is a chemical name?
 - i. Examples
 - c. What is a chemical symbol?
 - i. Examples
 - d. Atomic Mass (*See Lesson 1*)
4. What are the 3 types of elements?
 - a. Where can these elements be found on the Periodic Table?
 - b. What are the characteristics of these types of elements?
5. What is a group?
 - a. What do the elements in a group have in common?
6. What is a period?
 - a. What do the elements in a period have in common?

Lesson 3: Electrons and Chemical Bonding

1. What is a chemical bond?
 - a. Why do atoms form bonds?
 - i. How can you figure out what an atom will do to become "happy"?
2. What is a group of atoms that are chemically bonded called?
3. What are valence electrons?

Atoms Periodic Table Study Guide Answer

Leo J. Malone, Theodore O. Dolter



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