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web apr 7 2014 molarity moles of solute liters of solution 8 4 2 2 a first convert 250 ml to liters 250 1000 0 25 then calculate molarity 5 moles 0 25 liters 20 m 3 c a solution with molarity 2 requires 2 m of n a oh per liter so 4 x 2 8 m 4 a a solution of molarity 1 5 m requires 1 5 mol of na to every litre of solvent

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$$m_{\text{mix}} = \frac{m_1 v_1 + m_2 v_2}{v_1 + v_2}$$

$$\frac{1.5 \times 480 + 1.2 \times 520}{480 + 520} = \frac{720 + 624}{1000} = 1.344 \text{ m}$$
 q5 with increase in temperature which of these changes a molality b fraction of solute present in water

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