## Hess's Law Worksheet

 Calculate ΔH for the reaction: C<sub>2</sub>H<sub>4</sub> (g) + H<sub>2</sub> (g)→C<sub>2</sub>H<sub>5</sub> (g), from the following data.

 $\Delta H = -1411. \text{ kJ}$ 

$$C_2H_6(g) + 3\% O_2(g) \rightarrow 2 CO_2(g) + 3 H_2O(l)$$

 $\Delta H = -1560$ . kJ

$$H_2(g) + \% O_2(g) \rightarrow H_2O(l)$$

 $\Delta H = -285.8 \text{ kJ}$ 

 Calculate ΔH for the reaction 4 NH<sub>3</sub> (g) + 5 O<sub>2</sub> (g) EB4 NO (g) + 6 H<sub>2</sub>O (g), from the following data.

$$N_2(g) + O_2(g) \rightarrow 2 NO(g)$$

 $\Delta H = -180.5 \text{ kJ}$ 

$$N_2(g) + 3 H_2(g) \rightarrow 2 NH_3(g)$$

 $\Delta H = -91.8 \text{ kJ}$ 

 $\Delta H = -483.6 \text{ kJ}$ 

Find ΔH for the reaction 2H<sub>2</sub>(g) + 2C(s) + O<sub>2</sub>(g) → C<sub>2</sub>H<sub>5</sub>OH(I), using the following thermochemical data.

 $\Delta H = -875. \text{ kJ}$ 

$$C(s) + O_2(g) \rightarrow CO_2(g)$$

 $\Delta H = -394.51 \, kJ$ 

 $\Delta H = -285.8 \text{ kJ}$ 

Calculate 
 ☐H for the reaction CH<sub>4</sub> (g) + NH<sub>3</sub> (g) → HCN (g) + 3 H<sub>2</sub> (g), given:

$$N_2(g) + 3 H_2(g) \rightarrow 2 NH_3(g)$$

 $\Delta H = -91.8 \text{ kJ}$ 

 $\Delta H = -74.9 \text{ kJ}$ 

$$H_2(g) + 2C(s) + N_2(g) \rightarrow 2HCN(g)$$

 $\Delta H = +270.3 \text{ kJ}$ 

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**Heather Hattori, Richard H. Langley** 

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