

Acids & Bases HOMEWORK

Lesson 1

1. Provide the name for each of the strong acids listed below. Provide the formula for each of the strong bases listed below.

- ~~HBr~~ **hydrobromic acid**
- ~~HCl~~ **hydrochloric acid**
- ~~HNO₃~~ **nitric acid**
- ~~H₂SO₄~~ **sulfuric acid**
- ~~Strontium hydroxide~~ **LiOH**
- ~~potassium hydroxide~~ **KOH**
- ~~Sodium hydroxide~~ **NaOH**
- ~~calcium hydroxide~~ **Ca(OH)₂**
- ~~ammonium hydroxide~~ **LiOH**

H₂O₂ is a WEAK acid & does not completely dissociate in water.

2. Of the acids and bases listed below, circle the ones that completely dissociate in water.

HCl

~~H₂O~~

~~H₂SO₄~~

~~Mg(OH)₂~~

HNO₃

NaOH

~~NaOH~~

LiOH

3. What is the formula for hydronium? **H₃O⁺**

4. BaCu_2O_6 will react with water and lose a H^+ ion in the process.

- Write a balanced equation for this reaction: **$\text{BaCu}_2\text{O}_6 + \text{H}_2\text{O} \rightarrow \text{BaCu}_2\text{O}_5 + \text{H}^+$**
- Would you expect a large or small equilibrium constant for this reaction? **small, because BaCu_2O_6 is a weak acid.**
- Would you define BaCu_2O_6 as a proton donor or a proton acceptor? **proton donor, because it is losing a H^+ ion.**
- In your written equation, is BaCu_2O_6 acting as a Brønsted-Lowry acid or a Brønsted-Lowry base? **Brønsted-Lowry acid, because it is losing a proton.**
- Identify the conjugate base of BaCu_2O_6 : **BaCu_2O_5**
- Identify BaCu_2O_6 as polyprotic, amphoteric, or neither. **polyprotic, because it contains multiple H's.**

5. F^- will react with water and gain a H^+ ion in the process.

- Write a balanced equation for this reaction: **$\text{F}^- + \text{H}_2\text{O} \rightarrow \text{HF} + \text{OH}^-$**
- Would you expect a large or small equilibrium constant for this reaction? **small, because F^- is a weak base.**
- Would you define F^- as a proton donor or a proton acceptor? **proton acceptor, because it is gaining a H^+ ion.**
- In your written equation, is F^- acting as a Brønsted-Lowry acid or a Brønsted-Lowry base? **Brønsted-Lowry base, because it is gaining a proton.**
- Identify the conjugate acid of F^- : **HF**
- Identify F^- as polyprotic, amphoteric, or neither. **neither**

6. HPO_4^{2-} will react with water and lose a H^+ ion in the process.

- Write a balanced equation for this reaction: **$\text{HPO}_4^{2-} + \text{H}_2\text{O} \rightarrow \text{H}_2\text{PO}_4^- + \text{OH}^-$**
- Would you expect a large or small equilibrium constant for this reaction? **small, because HPO_4^{2-} is a weak acid.**
- Would you define HPO_4^{2-} as a proton donor or a proton acceptor? **proton donor, because it is losing a H^+ ion.**
- In your written equation, is HPO_4^{2-} acting as a Brønsted-Lowry acid or a Brønsted-Lowry base? **Brønsted-Lowry acid, because it is losing a proton.**
- Identify the conjugate base of HPO_4^{2-} : **PO_4^{3-}**
- Identify HPO_4^{2-} as polyprotic, amphoteric, or neither. **amphoteric, because it has a negative charge (and can take a proton) and it has H's (and can lose one).**

19 3 Acids And Bases Answer Key Chemistry

N Colangelo



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High Temperature Polymer Electrolyte Membrane Fuel Cells Qingfeng Li, David Aili, Hans Aage Hjuler, Jens Oluf Jensen, 2015-10-15 This book is a comprehensive review of high temperature polymer electrolyte membrane fuel cells PEMFCs PEMFCs are the preferred fuel cells for a variety of applications such as automobiles cogeneration of heat and power units emergency power and portable electronics The first 5 chapters of the book describe rationalization and illustration of approaches to high temperature PEM systems Chapters 6-13 are devoted to fabrication optimization and characterization of phosphoric acid doped polybenzimidazole membranes the very first electrolyte system that has demonstrated the concept of and motivated extensive research activity in the field The last 11 chapters summarize the state of the art of technological development of high temperature PEMFCs based on acid doped PBI membranes including catalysts electrodes MEAs bipolar plates modelling stacking diagnostics and applications

The American chemist ,1874

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