

## 19.2

## HYDROGEN IONS AND ACIDITY

## Section Review

## Objectives

1. Calculate  $K_a$  and  $K_b$  for weak acids and bases, and  $pK_a$  and  $pK_b$  for weak acids and bases.
2. Calculate  $K_a$  and  $K_b$  for weak acids and bases, and  $pK_a$  and  $pK_b$  for weak acids and bases.
3. Calculate the  $pH$  of weak acid and base solutions.

## Vocabulary

- |                 |                                      |                 |
|-----------------|--------------------------------------|-----------------|
| acid            | acid dissociation constant ( $K_a$ ) | acid ionization |
| acidic solution | acidic solution                      | acidic solution |
| acidic solution | acidic solution                      | acidic solution |

## Key Equations

- |   |  |
|---|--|
| $K_a = \frac{[H^+][A^-]}{[HA]}$ (Equation 19.1) | $K_b = \frac{[OH^-][HB^+]}{[B]}$ (Equation 19.2) |
|---|--|

## Part A Completion

Write the appropriate formula or name for each compound, and write its  $K_a$  or  $K_b$  value.

1.  $K_a$  of acetic acid,  $CH_3COOH$ , is  $1.8 \times 10^{-5}$ .

2.  $K_b$  of ammonia,  $NH_3$ , is  $1.8 \times 10^{-5}$ .

3.  $K_a$  of hydrofluoric acid,  $HF$ , is  $6.6 \times 10^{-4}$ .

4.  $K_b$  of methylamine,  $CH_3NH_2$ , is  $4.4 \times 10^{-4}$ .

5.  $K_a$  of formic acid,  $HCOOH$ , is  $1.8 \times 10^{-4}$ .

6.  $K_b$  of pyridine,  $C_5H_5N$ , is  $1.7 \times 10^{-9}$ .

7.  $K_a$  of benzoic acid,  $C_6H_5COOH$ , is  $6.3 \times 10^{-5}$ .

8.  $K_b$  of triethylamine,  $(C_2H_5)_3N$ , is  $5.9 \times 10^{-4}$ .

9.  $K_a$  of lactic acid,  $CH_3CH(OH)COOH$ , is  $1.4 \times 10^{-4}$ .

10.  $K_b$  of diethylamine,  $(C_2H_5)_2NH$ , is  $1.3 \times 10^{-3}$ .

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# 19 2 Hydrogen Ions And Acidity Section Review Answers

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