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web name the following molecular compounds scl2 n2o4 p4o10 pf5 scl2 sulfur dichloride n2o4 dinitrogen tetroxide the a of tetra is dropped before the o of oxide p4o10 tetraphosphorous decoxide the a of deca is dropped before the o of oxide pf5

phosphorous pentafluoride

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web one of the ions produced is always H^+ which immediately combines with a water molecule to form the hydronium ion H_3O^+ the H^+ in is what defines the acidic properties of a substance because of the special classification acids have a naming system different from ionic or other molecular covalent compounds

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web examine the pairs of ternary acids in model 2 that contain sulfur phosphorus and nitrogen each pair has one acid that ends in ic and another that ends in ous these endings are related to the name of the polyatomic anion found in the acid ate or ite

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web the acids lose a hydrogen ion for each acid base reaction in model 2 describe the role of the brønsted lowry base in the proton H^+ ion transfer that occurs the base gains or accepts a hydrogen ion proton as you saw in last week's pogil all arrhenius bases in model 1 have an OH^- ion in their chemical formulas

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web the molecular or empirical formula acids have their own set of naming rules cf sec 2 8 which you must also know you also need to be able to recognize and name a few organic compounds such as simple alkanes and alcohols see sec 2 9 key questions 8 in the chemical formula and name which element is given first a metal or nonmetal 9

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