



Alero Manual

Wikipedia contributors



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web high hazard testing the big explosives experimental facility beef is a high explosive testing facility providing data to

support the stockpile stewardship program and other national security programs it is just one of the several high hazard facilities at the nns threat detection crne

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web nonproliferation test and evaluation complex southern nevada including the site is a dry climate and summer temperatures can top 100 degrees while winter can bring snow bring plenty of water and hydrate before and during the tour you can bring snacks lunch and non alcoholic beverages on the bus

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web the nevada test site is an area set aside for the testing of nuclear weapons it is looked after by the united states department of energy it is in nye county nevada about 65 miles 105 km northwest of las vegas the site was begun on january 11 1951 it covers an area of 1 350 square miles 3 500 km² of desert and mountain lands

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web to understand exactly what was accomplished in terms of safely testing hundreds of nuclear devices atomic bombs during the active times at the nevada test site you really need to take an on site tour if you can arrange it failing that a visit to the atomic test museum the yucca mountain museum and the nts history center might suffice

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 $2 \text{ CO}_2(\text{g}) + \text{H}_2\text{O}(\text{g}) \rightarrow \text{C}_2\text{H}_2(\text{g}) + \frac{5}{2} \text{O}_2(\text{g})$ $\Delta h = 94.5 \text{ kJ}$
 $\text{H}_2\text{O}(\text{g}) + \frac{1}{2} \text{O}_2(\text{g}) \rightarrow \text{H}_2\text{O}(\text{l})$ $\Delta h = 71.2 \text{ kJ}$
 $\text{C}_2\text{H}_6(\text{g}) + \frac{7}{2} \text{O}_2(\text{g}) \rightarrow 2 \text{ CO}_2(\text{g}) + 3 \text{ H}_2\text{O}(\text{g})$ $\Delta h = 283 \text{ kJ}$

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 $\text{C}_2\text{H}_4(\text{g}) + 3 \text{ O}_2(\text{g}) \rightarrow 2 \text{ CO}_2(\text{g}) + 2 \text{ H}_2\text{O}(\text{l})$ $\Delta h = 1411 \text{ kJ/mole}$
 $\text{C}_2\text{H}_6(\text{g}) + \frac{7}{2} \text{ O}_2(\text{g}) \rightarrow 2 \text{ CO}_2(\text{g}) + 3 \text{ H}_2\text{O}(\text{l})$ $\Delta h = 1560 \text{ kJ/mole}$
 $\text{H}_2\text{O}(\text{g}) \rightarrow \text{H}_2\text{O}(\text{l})$ $\Delta h = 285 \text{ kJ/mole}$

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process is the sum of the Δh values for each step to use hess s law two principles must be understood one if an equation is reversed the sign of the Δh value is also reversed two if an equation is multiplied by a

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web may 8 2015 hess law lab by maya parks partners ben seufert kelsea floyd convert joules to kj in your final answer multiply the mass by the change in temperature and the c given trial 1 use hess s law and the following equations and Δh values to determine the heat of reaction for the reaction c 2 h 4

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web thus for reaction 9 2 in which 1 mole of h₃o and 1 mole of oh combine to form 2 moles of h₂o h 55 8 kj as shown note that the enthalpy change h hproducts hreactants is positive if heat is absorbed that is if hproducts

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web objective the purpose of the lab is to combine two reactants in the calorimeter and measure the heats of reaction in solution eventually allowing an individual to prove or disprove hess s law introduction thermodynamics is the study of energy exchanged between the system and the surroundings

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