Name _____ *Key*_____

1. Hydrogen, H_2, is used as rocket fuel. The hydrogen is burned in oxygen, O_2, according to the following equation. What is the enthalpy change if 200.0 g of hydrogen, H_2 is burned?

$$2\ H_2(g)\ +\ O_2(g)\ \rightarrow\ 2\ H_2O(g)\ ;\ \Delta H = -484\ kJ$$

$$\frac{200.0\ g\ H_2}{}\ \left|\ \frac{1\ mol\ H_2}{2.016\ g\ H_2}\right.\ =\ 99.2\ mol\ H_2 \qquad \frac{99.2\ mol\ H_2}{}\ \left|\ \frac{-484\ kJ}{2\ mol\ H_2}\right.$$

$$\boxed{-24,006\ kJ}$$

2. Carbon disulfide burns in air, producing carbon dioxide and sulfur dioxide.

$$CS_2(l)\ +\ 3\ O_2(g)\ \rightarrow\ CO_2(g)\ +\ 2\ SO_2(g)\ ;\ \Delta H = -1077\ kJ$$

What is ΔH for the following reaction?

reverse so change sign of
ΔH - cut in ½
quantities in ½

$$\tfrac{1}{2}\ CO_2(g)\ +\ SO_2(g)\ \rightarrow\ \tfrac{1}{2}\ CS_2(l)\ +\ \tfrac{3}{2}\ O_2(g)$$

$$\frac{+1077\ kJ}{2}\ =\ \boxed{538.5\ kJ}$$
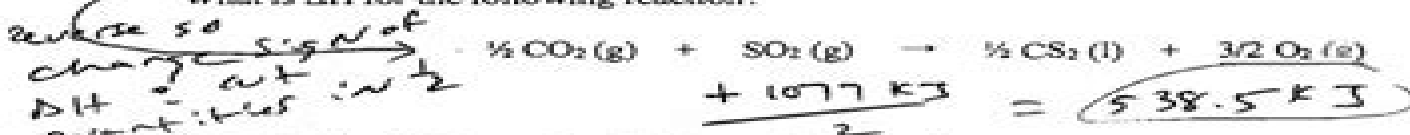
3. Following is the equation for the combustion of acetone:

$$2\ C_3H_6O(g)\ +\ 8\ O_2(g)\ \rightarrow\ 6\ H_2O(g)\ +\ 6\ CO_2(g)\ +\ 853.6\ kcal$$

Exothermic $/\ 2\ mol\ C_3H_6O$

How much heat will be given off (in kcal) if 4.5 mol of acetone is burned completely?

$$\frac{-853.6\ kcal}{2\ mol\ C_3H_6O}\ \left|\ 4.5\ mol\ C_3H_6O\right.\ =\ \boxed{-1920.6\ kcal}$$

4. The specific heat of iron is 0.110 cal / g-°C. How much heat must be added to 5.2 g of iron at 25 °C to raise its temperature to 225 °C?

$$heat = q = m\ c\ \Delta T\ =\ (5.2)\left(\frac{0.110\ cal}{g\ °C}\right)(225-25\ °C)\ =\ \boxed{114.4\ cal}$$

Endothermic

5. Hydrazine, N_2H_4, is a colorless liquid used as a rocket fuel. What is the enthalpy change for the process in which hydrazine is formed from its elements (reaction (1) below)? Use Hess's Law to calculate ΔH for reaction (1) given the data below.

$$(1)\qquad N_2(g)\ +\ 2H_2(g)\ \rightarrow\ N_2H_4(l)$$

$$N_2H_4(l)\ +\ O_2(g)\ \rightarrow\ N_2(g)\ +\ 2H_2O(l)\ ;\ \Delta H = -622.2\ kJ$$

$$H_2(g)\ +\ \tfrac{1}{2}\ O_2(g)\ \rightarrow\ H_2O(l)\ ;\ \Delta H = -285.8\ kJ$$

Reverse

$$N_2(g)\ +\ 2\ H_2O(l)\ \longrightarrow\ N_2H_4(l)\ +\ O_2(g)\qquad \Delta H = 622.2\ kJ$$

double quantity

$$2\ H_2(g)\ +\ O_2(g)\ \longrightarrow\ 2\ H_2O(l)\qquad \Delta H = (2)(-285.8)\ kJ$$

$$\overline{N_2(g)\ +\ 2\ H_2(g)\ \longrightarrow\ N_2H_4(l)}$$

$$\boxed{\Delta H = +50.6\ kJ}$$
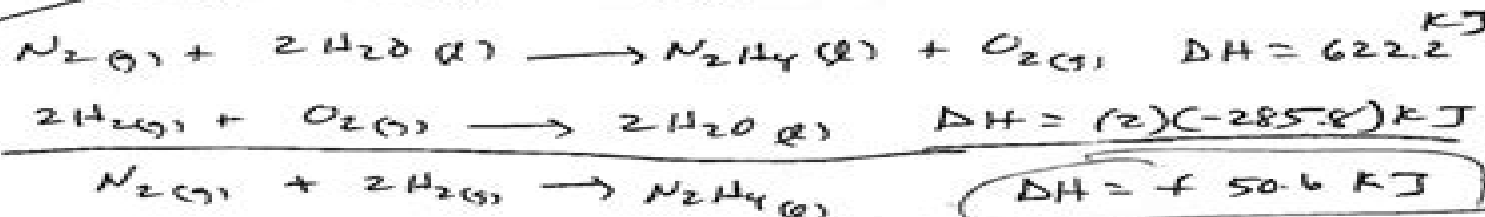
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Disha Experts

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