

of substance:

Substance:

of particles of a substance
in mol

6.02×10^{23} particles
amount of substance

Constant:

particles per mol, $6.02 \times 10^{23} \text{ mol}^{-1}$

Unit:

mol, measured in g
equivalent to mol

Atomic Structure

Isotopes:

→ Same # of protons, same element
of p^+ and e^- but different
neutrons different mass

Physical Properties:

→ Can be different - e.g. different
melting + boiling points
→ i.e. heavy water has higher melting
point + boiling point + density

Chemical Properties:

→ Same no. of electrons + same electronic
structure in outer shell + transfer of e^- is
unaffected hence have same chemical
properties.
→ Different no. of n has no effect
on bonding + reactions

Compounds + Ionic Formulas:

Alkali NO_3^-

Molar Gas Volume:

→ Volume of gas per mole of gas molecules in a stated temperature + pressure
- Measured in $\text{dm}^3 \text{ mol}^{-1}$ or $\text{m}^3 \text{ mol}^{-1}$

At RTP (20°C and 100kPa):
(293K and 100000Pa)

is $24\text{ dm}^3 \text{ mol}^{-1}$
at RTP.

Volume /
 dm^3

Electronegativity:

→ the measure of attraction of a bonded atom
in a covalent bond

Electronegativity increases across period:

- Increase in no. of protons in an atom
→ Nuclear charge increases

Electronegativity DE:

- More electron shells
→ More shielding
→ Atomic radius increases
→ Nuclear charge decreases
→ Nuclear attraction decreases
→ Electronegativity decreases

OCR A LEVEL CHEMISTRY MODULE 2 DIAGRAMS



→ Consists of a small dense
nucleus of protons + neutrons
surrounded by electrons orbiting
in electron shells in shells;
- No. of p^+ = No. of e^-

→ Equal + opposite charges cancel
in atom has neutral charge
- Neutrons important in nuclear
stability - Electrostatic repulsion
between p^+ in small dense nucleus

Mass Number:

→ No. of p^+ + n in nucleus of atom
- p^+ and n have relative mass of 1
and e^- mass is negligible is 0

Relative Isotopic Mass:

→ The mass of an atom of an isotope
relative to $1/12$ th the mass of an atom of ^{12}C

Relative Atomic Mass:

→ The weighted mean mass of an atom of
an element relative to $1/12$ th the mass of
an atom of ^{12}C

Weighted Mean Mass:

→ Average mass of all isotopes, taking into
account the relative abundance of each isotope

- If dipole
molecule is
out + molts

e.g. H_2O

- O is more electronegative than H
- O-H has polar bond
since do not cancel there is an overall dipole

e.g. CO_2

O is more electronegative than C
 $\text{C}=\text{O}$ has a dipole
but bond dipoles cancel

A2 Chemistry Ocr Rates Practical Exam Quantitative

Mike Smith, John Older



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NEW A-LEVEL CHEMISTRY FOR 2018 CGP. BOOKS, 2018 Revise A2 Chemistry for OCR A Lord Eccles, Wooster, 2005-03-29 With short questions at the end of each section that make students stop and think about the topic this work provides tips on common pitfalls and advice on how to tackle different types of exam question and exam preparation It also includes practice exam style questions

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data people and process the physical parts of computing devices those that you can actually touch are referred to as
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