

## 19.2

## HYDROGEN IONS AND ACIDITY

## Section Review

## Objectives

1. Calculate  $K_a$  and  $K_b$  for weak acids and bases, and  $pK_a$  and  $pK_b$  for weak acids and bases.
2. Calculate  $K_a$  and  $K_b$  for weak acids and bases, and  $pK_a$  and  $pK_b$  for weak acids and bases.
3. Calculate the  $pH$  of weak acid and base solutions.

## Vocabulary

- |                 |                                      |                 |
|-----------------|--------------------------------------|-----------------|
| acid            | acid dissociation constant ( $K_a$ ) | acid ionization |
| acidic solution | acidic solution                      | acidic solution |
| acidic solution | acidic solution                      | acidic solution |

## Key Equations

- |   |  |
|---|--|
| $K_a = \frac{[H^+][A^-]}{[HA]}$ (Equation 19.1) | $K_b = \frac{[OH^-][HB^+]}{[B]}$ (Equation 19.2) |
|---|--|

## Part A Completion

Write the appropriate formula or name for each compound, and write its  $K_a$  or  $K_b$  value.

1.  $K_a$  of  $HNO_3$  is  $2.4 \times 10^{-4}$ . Name the acid.

2.  $K_b$  of  $CH_3NH_2$  is  $4.4 \times 10^{-4}$ . Name the base.

3.  $K_a$  of  $HClO_4$  is  $1.0 \times 10^{-7}$ . Name the acid.

4.  $K_b$  of  $CH_3COO^-$  is  $5.6 \times 10^{-10}$ . Name the base.

5.  $K_a$  of  $H_2SO_4$  is  $1.0 \times 10^{-2}$ . Name the acid.

6.  $K_b$  of  $CH_3NH_2$  is  $4.4 \times 10^{-4}$ . Name the base.

7.  $K_a$  of  $HNO_3$  is  $2.4 \times 10^{-4}$ . Name the acid.

8.  $K_b$  of  $CH_3NH_2$  is  $4.4 \times 10^{-4}$ . Name the base.

9.  $K_a$  of  $HClO_4$  is  $1.0 \times 10^{-7}$ . Name the acid.

10.  $K_b$  of  $CH_3COO^-$  is  $5.6 \times 10^{-10}$ . Name the base.

11.  $K_a$  of  $H_2SO_4$  is  $1.0 \times 10^{-2}$ . Name the acid.

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# 19 2 Hydrogen Ions And Acidity Section Review Answers

**Kevin T. Patton, Frank B. Bell, Terry  
Thompson, Peggie L. Williamson**

## 19 2 Hydrogen Ions And Acidity Section Review Answers:

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