Chemical Equations and Reactions

SECTION 1: Describing Chemical reactions

Match the symbol on the left with its appropriate description on the right.

(b) A gas forms.

(c) A reversible reaction occurs.

(a) A precipitate forms.

(d) Heat is applied to the reactants.

(e) A chemical is dissolved in water.

- (f) A chemical is in the liquid state.
- 2. Finish balancing the following equation:

 $3Fe_3O_4 + Al \longrightarrow Al_2O_3 + Fe$

3. In each of the following formulas, write the total number of atoms present.

a. 4SO₂

c. 3Al₂(SO₄)₃

4. Convert the following word equation into a balanced chemical equation: aluminum metal + copper(II) fluoride -> aluminum fluoride + copper metal

4Al + 2CuF₂ → 2AlF₃ + 2Cu

$$2AI + 3CuF_2 \longrightarrow 2AIF_3 + 3Cu$$

 $A1 + 3CuF_2 \longrightarrow AIF_4 + 3Cu$

5. One way to test the salinity of a water sample is to add a few drops of silver nitrate solution with a known concentration. As the solutions of sodium chloride and silver nitrate mix, a precipitate of silver chloride forms, and sodium nitrate is left in solution. Translate these sentences into a balanced chemical equation.

 $NaCl(aq) + AgNOy(aq) \longrightarrow AgCl(s) + NaNOy(aq)$

Describing Chemical Equations 11 1 Key

Jessica J Manson

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